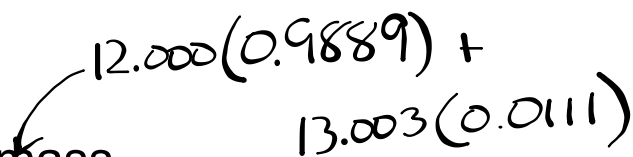


Unit 1 - Classification of Matter

- Types of matter
- Periodic table - trends, families, etc.
- Periodic law
- Elements
- Atoms
- Isotopes
- Calculating atomic mass
- Ions
- Bohr Theory
- Quantum Mechanical Model
- Electron configurations

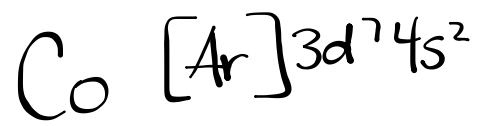
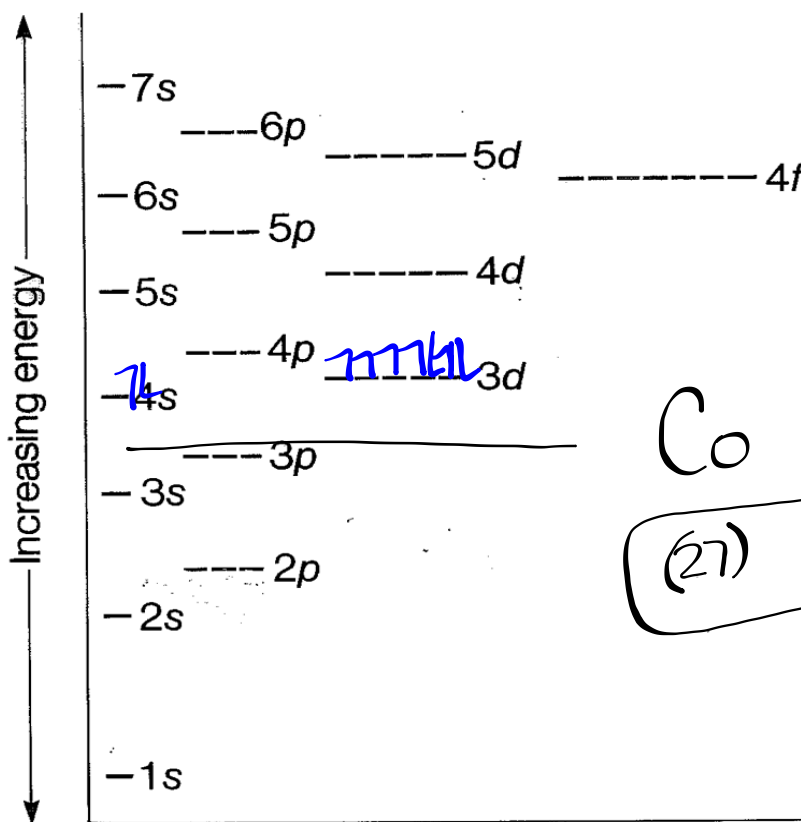
$$12.000(0.9889) + 13.003(0.0111)$$
A handwritten calculation in black ink showing the weighted average of two isotopes. The first term is 12.000 multiplied by 0.9889, and the second term is 13.003 multiplied by 0.0111. A curved arrow points from the first term to the bullet point 'Calculating atomic mass' in the list above.



Isotope Name	Atomic Number	Mass Number	Symbol	# of Protons	# of Neutrons

Name	Symbol	# of Protons	# of Electrons	Gain or Lose?	Net Charge
Selenide ion	Se^{2-}				

Aufbau Diagram



(27) 18 + 9

Test Review

Ch. 4 p. 122-123 #34, 39-41, 44-58, 60, 63-65,
71-73

Ch. 5 p. 149-150 #23-39

Ch. 6 p. 181-182 #24, 26-35