Unit 1 - Classification of Matter

- Types of matter
- Periodic table - trends, families, etc.
- Periodic law
- Elements
- Atoms
- Isotopes
- Calculating atomic mass
- Ions
- Bohr Theory
- Quantum Mechanical Model
- Electron configurations

12.000(0.9889) +
13.003(0.011)

$^{12}_{6}C$

$6p^+$

$6e^-$

$6n$

$^{13}_{6}C$

$6p^+$

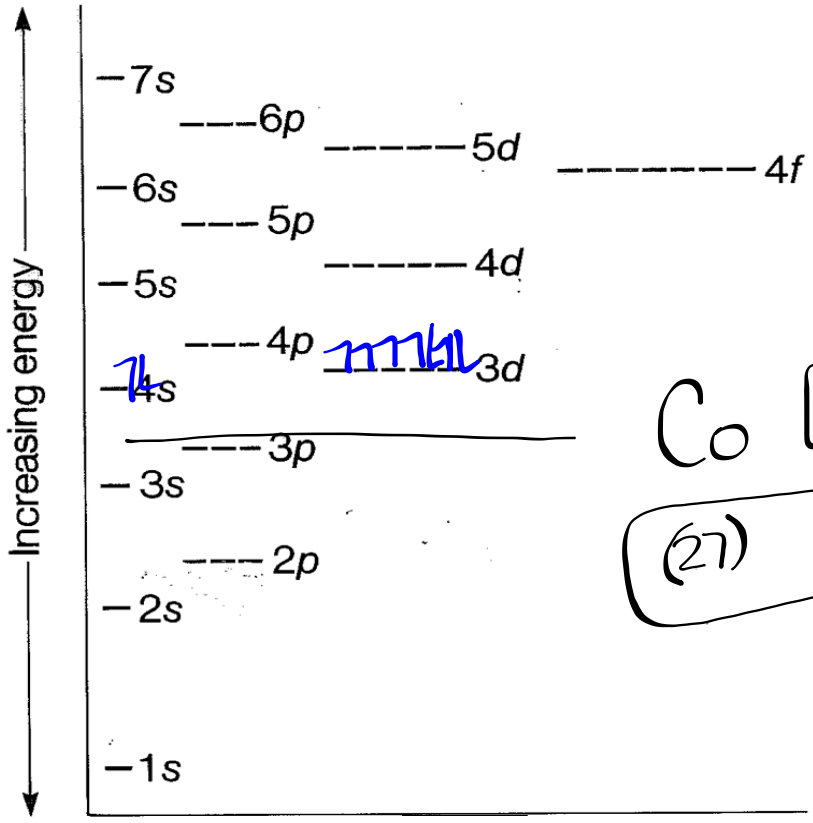
$6e^-$

$7n$

Isotope Name	Atomic Number	Mass Number	Symbol	# of Protons	# of Neutrons

Name	Symbol	# of Protons	# of Electrons	Gain or Lose?	Net Charge
Selenide ion	Se^{2-}				

Aufbau Diagram



Co [Ar] $3d^7 4s^2$

(27) 18 + 9

Test Review

Ch. 4 p. 122-123 #34, 39-41, 44-58, 60, 63-6 5, 71-73

Ch. 5 p. 149-150 #23-39

Ch. 6 p. 181-182 #24, 26-35