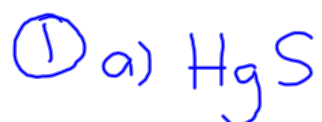


# Homework - Worksheet #1,3



③ a) copper(II) sulfide

b) copper(I) sulfide

c) magnesium oxide

d) aluminum oxide

e) lead(IV) sulfide

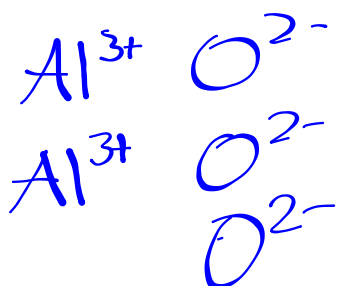
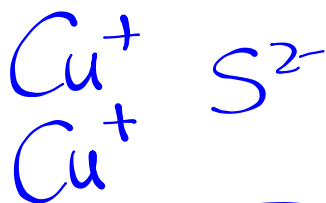
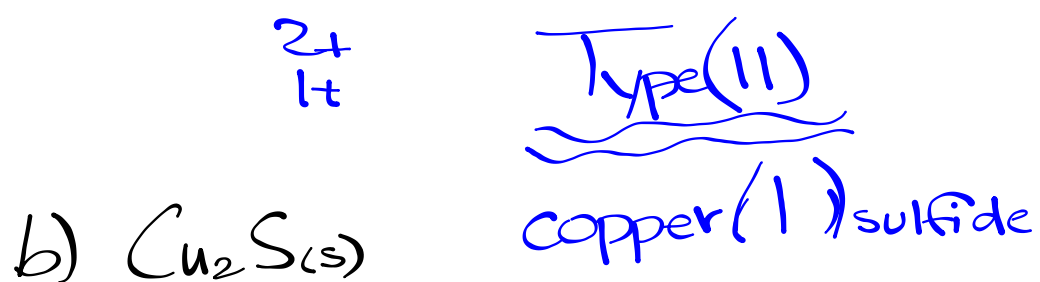
f) iron(III) oxide

g) zinc sulfide

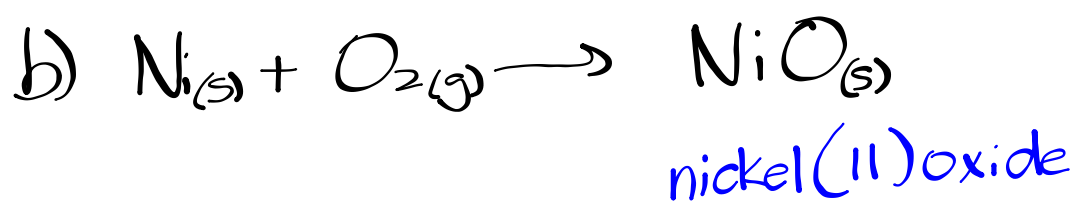
h) chromium(III) oxide

i) calcium chloride

j) vanadium(V) oxide



aluminum oxide



# Worksheet #2

## Writing Formulas for Polyatomic Ions

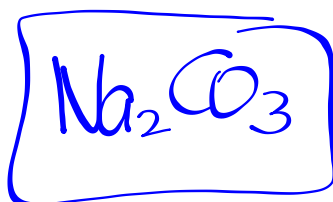
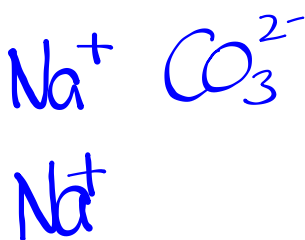
polyatomic ions- are ions made up of more than one atom.

Ex.  $\text{CO}_3^{2-}$  carbonate ion

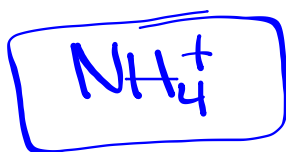
All polyatomic ions can be found with their charges on the back cover of the book.

Determining formulas of polyatomic ion compounds are done in the same manner as binary compounds.

Ex. sodium carbonate



Ex. calcium nitrate



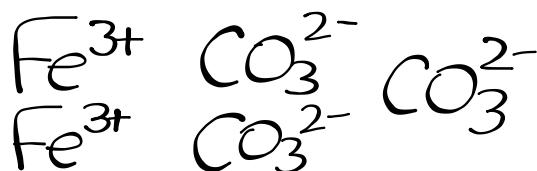
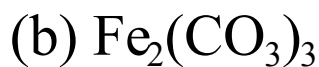
Calcium sulfate

- ate
- ite
- ide

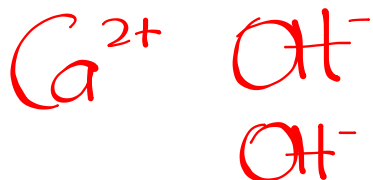
Name the following polyatomic compounds:



Sodium bicarbonate



iron (III) carbonate





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Sheet - Ionic Compounds

I	II										III	IV	V	VI	VII	VIII	
H																He	
Li	Be										B	C	N	O	F	Ne	
Na	Mg										Al	Si	P	S	Cl	Ar	
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Cs	Ba		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra		Rf	Db	Sg	Bh	Hs	Mt	Uun	Uuu	Uub						
			La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu
			Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

