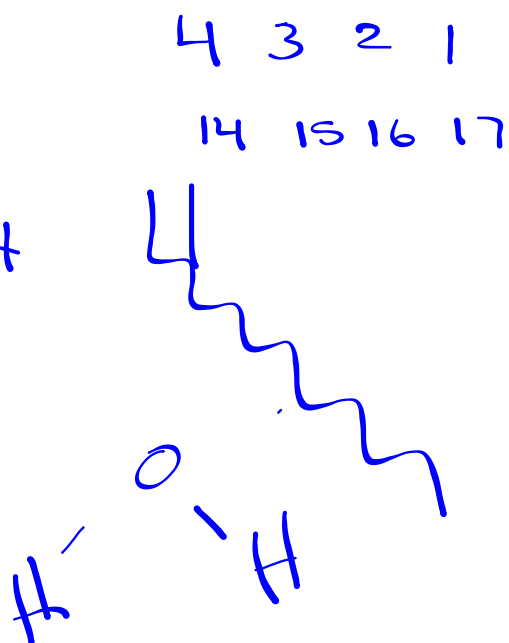
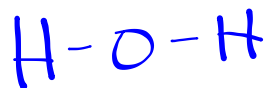
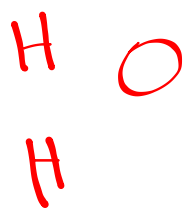
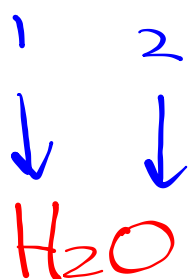


Unit 2 - Compounds

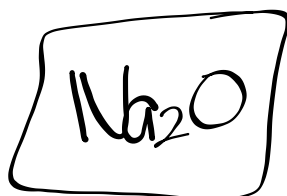
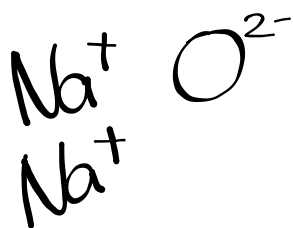
- Properties of Ionic Compounds, Molecular Compounds, Acids, and Bases (Empirical and Theoretical)
- Naming Ionic Compounds
- Writing formulas for Ionic Compounds
- Ionic hydrates
- Naming Molecular Compounds
- Writing formulas for Molecular Compounds
- Molecular Elements (Br HONCl IF) P₄ S₈
- Drawing structural diagrams
- Naming and writing formulas for Acids and Bases
- Lab - Identifying Unknown Compounds

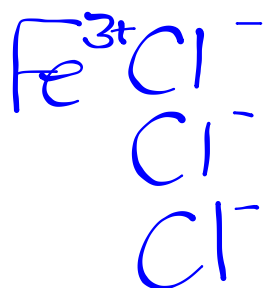




calcium chloride

sodium oxide





iron (III) chloride

copper (II) sulfide

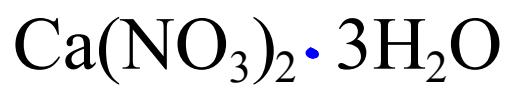




lithium nitrate

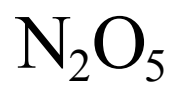
sodium sulfate





Calcium nitrate trihydrate

" " - 3-water



molecular

dinitrogen pentoxide

HClO

$H^+ ClO^-$ aqueous hydrogen
hypochlorite
hypochlorous acid

-ate
chromic acid

$H^+ CrO_4^{2-}$
 H^+



Review Questions p. 281-282

Worksheets

43-61, 65-71

