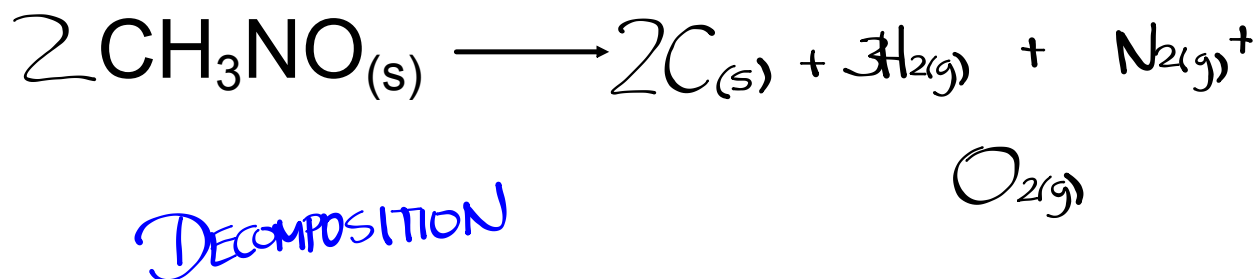
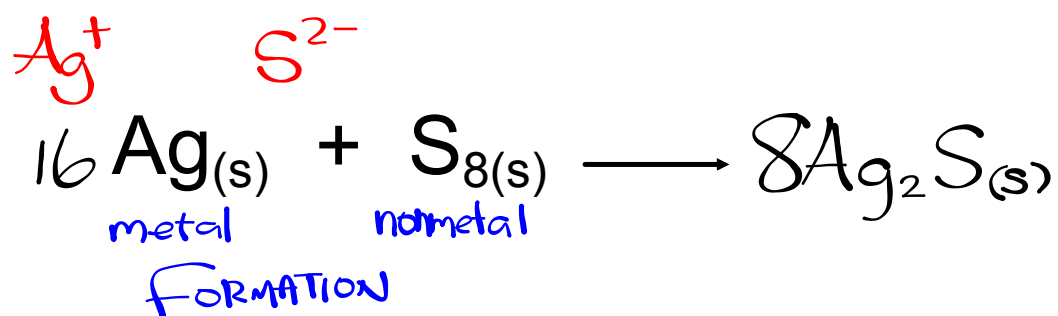


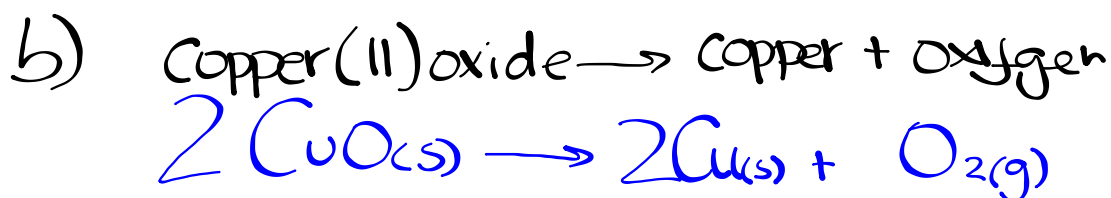
Warm Up



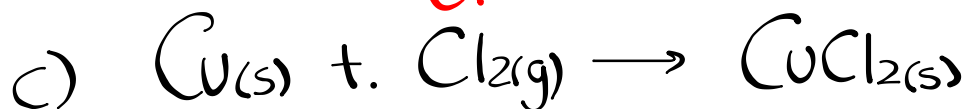
Check Homework - Worksheet



DECOMP.



FORMATION



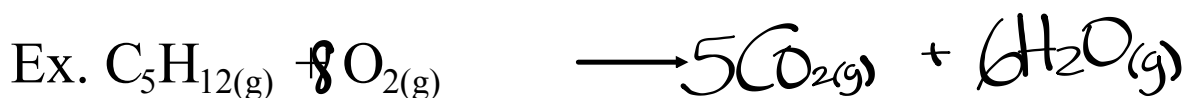
Chemical Reactions

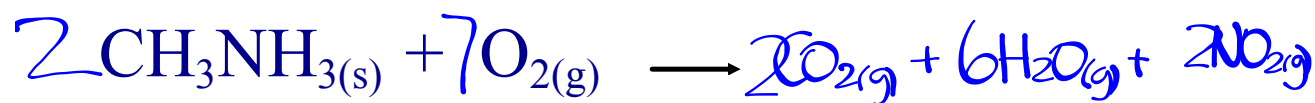
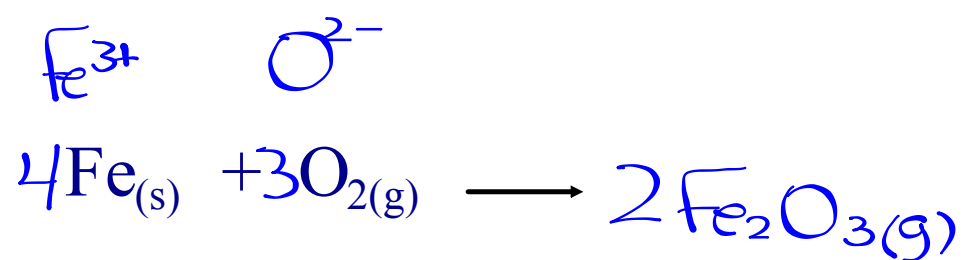
III. Combustion Reaction

A complete combustion reaction is the burning of a substance with oxygen to produce the most common oxides of the elements in the substance being burned.

Most Common Oxides:

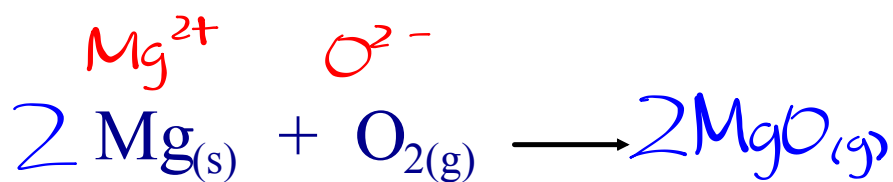
- Carbon : $\text{CO}_{2(g)}$
- Hydrogen: $\text{H}_2\text{O}_{(g)}$
- Sulfur: $\text{SO}_{2(g)}$
- Nitrogen: $\text{NO}_{2(g)}$
- A metal: Oxide of metal with most common ion charge





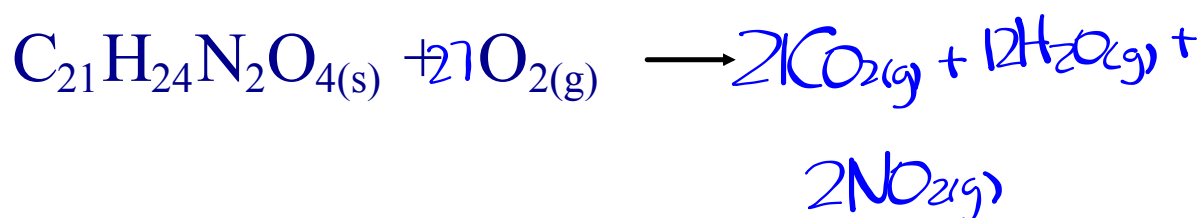
Combustion Reactions

Write a balanced chemical equation for the following combustion reactions:



4

54



Homework

p. 331 #13, 14

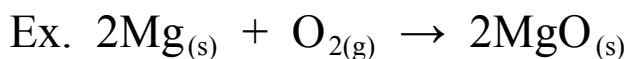
p. 332 #15, 16

p. 337 #20, 21

Chemical Reactions

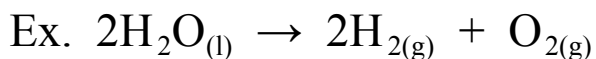
I. Formation Reactions

elements compound



II. Decomposition Reactions

compound elements



III. Combustion Reaction

substance + ^{O_{2(g)}}oxygen \longrightarrow most common oxides



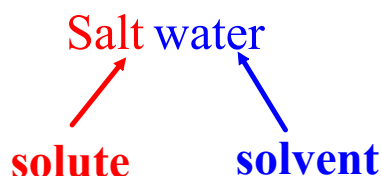
Chemical Reactions in Solution

Solution - homogeneous (uniform) mixture of a solute and a solvent.

⇒ solute - substance dissolved

⇒ solvent - substance doing dissolving (liquid)

Ex.



If the amount of solute that can dissolve in a solvent is large, then the solute is said to have *ahigh solubility*.

If the amount of solute that can dissolve in a solvent is small, then the solute is said to have *alow solubility*.

Solid substances formed from reactions in solutions are known as **precipitates**.