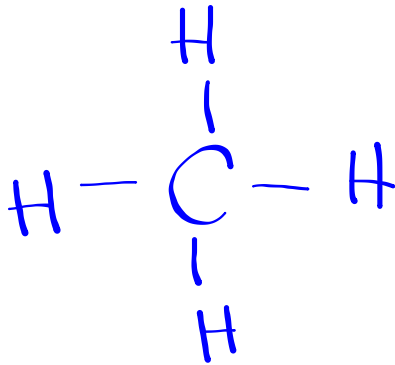


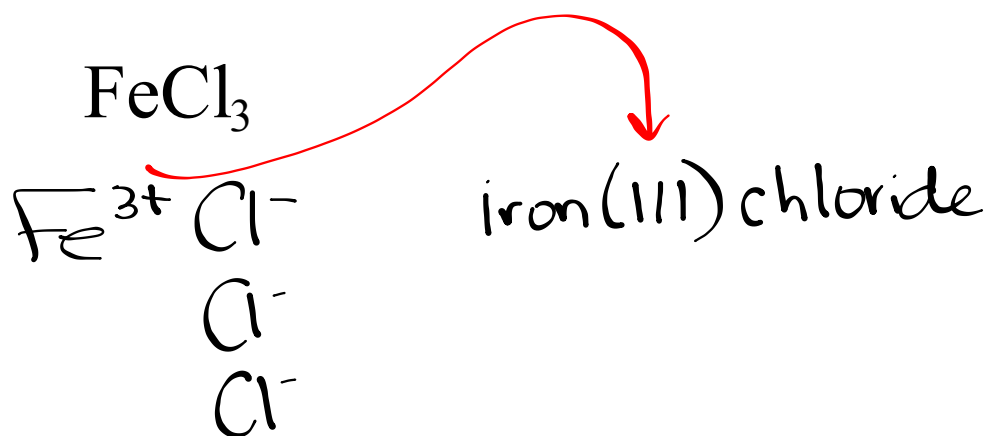
## Unit 2 - Compounds

- Properties of Ionic Compounds, Molecular Compounds, Acids, and Bases (Empirical and Theoretical)
- Naming Ionic Compounds
- Writing formulas for Ionic Compounds
- Ionic hydrates
- Naming Molecular Compounds
- Writing formulas for Molecular Compounds
- Molecular Elements → Br HONClIF (P<sub>4</sub>, S<sub>8</sub>)
- Drawing structural diagrams CH<sub>4</sub>
- Naming and writing formulas for Acids and Bases
- Lab - Identifying Unknown Compounds

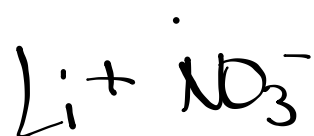




sodium oxide

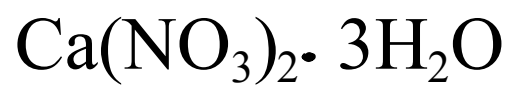


copper (II) sulfide



lithium nitrate

sodium sulfate

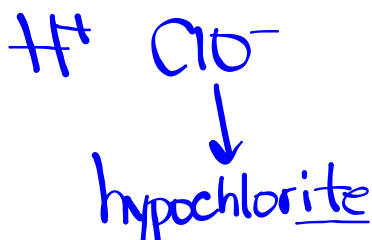


Calcium nitrate trihydrate  
" " - 3-water



dinitrogen pentoxide

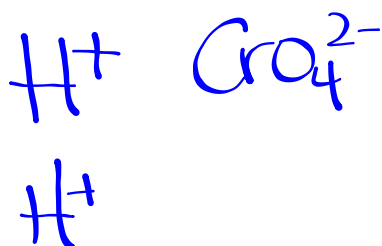
molecular



aqueous hydrogen hypochlorite  
or

hypochlorous acid

"ate"  
chromic acid





**Review Questions p. 281-282**

# **Worksheet**

**# 43-61, 65-71**

51, 53, 57- 61,65-71