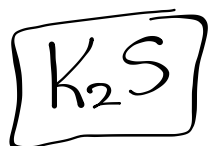
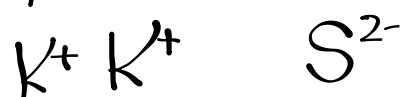
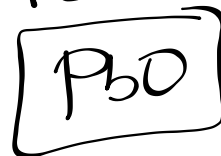


## Homework - Binary Ionic Compounds Sheet

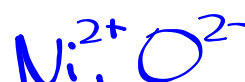
1. d) potassium sulfide



o) lead(II) oxide

2. b)  $AlCl_3$ 

aluminum chloride

o)  $NiO$ 

nickel(II)oxide

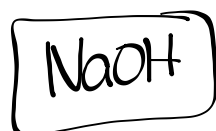
# Polyatomic Compounds

many atoms

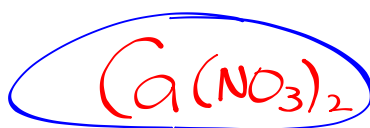
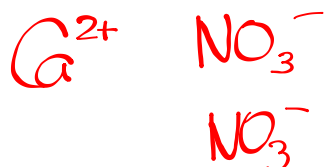
## Polyatomic Compounds

- polyatomic ions are groups of atoms that have a single charge  
Ex.  $\text{NO}_3^-$  (nitrate ion)
- Table 2, pg. 196
- they are to be balanced just like any other ion
- the name does not change to -ide

Ex. sodium + hydroxide = sodium hydroxide



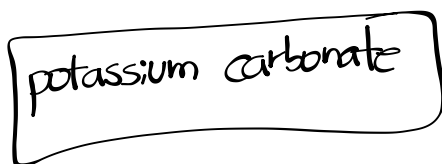
Calcium nitrate



## Naming Polyatomic Ions

Combination of the name of the metal, and the name of the polyatomic ion.

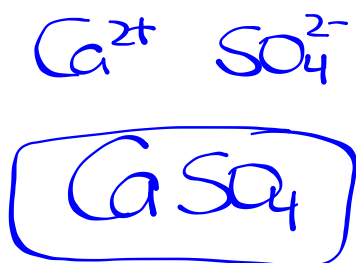
Ex.  $\text{K}_2\text{CO}_3$



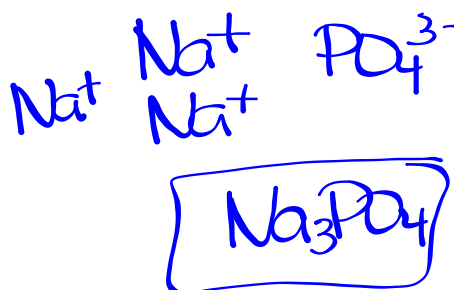
Calcium

Write the chemical formula of the following ionic compounds:

(a) calcium sulfate



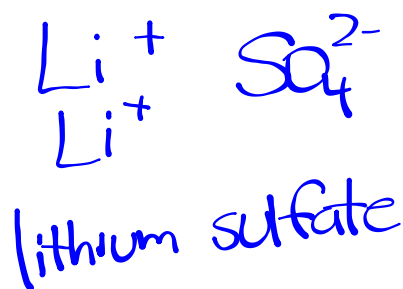
(b) sodium phosphate



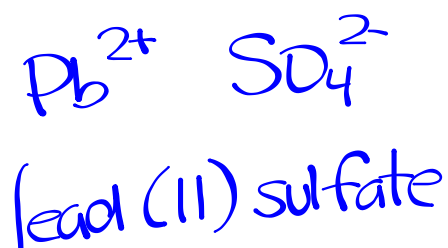
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Write the name of the following ionic compounds:

(a)  $\text{Li}_2\text{SO}_4$



(b)  $\text{PbSO}_4$



**p.198 #1-7  
(omit #5)**