

Homework p. 256/258

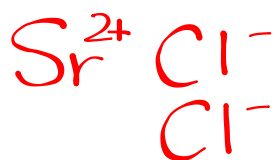


Handwritten annotations on the periodic table:

- Group 1: $1+$
- Group 2: $2+$
- Group 13: $3+$
- Group 17: $1-$
- Element Cl (Chlorine) is circled in black.
- Red arrows point from the '2+' group to the lanthanide and actinide series.

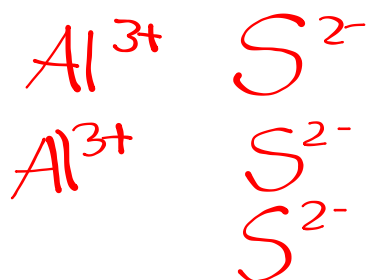
I	II										III	IV	V	VI	VII	VIII	
H																He	
Li	Be										B	C	N	O	F	Ne	
Na	Mg										Al	Si	P	S	Cl	Ar	
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
Cs	Ba		Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
Fr	Ra		Rf	Db	Sg	Bh	Hs	Mt	Uun	Uuu	Uub						
			La	Ce	Pr	Nd	Pm	Sm	Eu	Gd	Tb	Dy	Ho	Er	Tm	Yb	Lu
			Ac	Th	Pa	U	Np	Pu	Am	Cm	Bk	Cf	Es	Fm	Md	No	Lr

Type I Binary Ionic Compounds



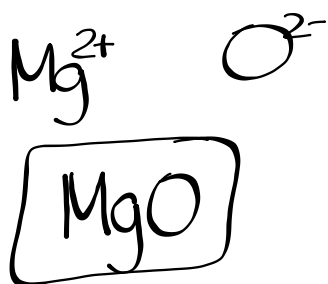
strontium chloride

aluminum sulfide

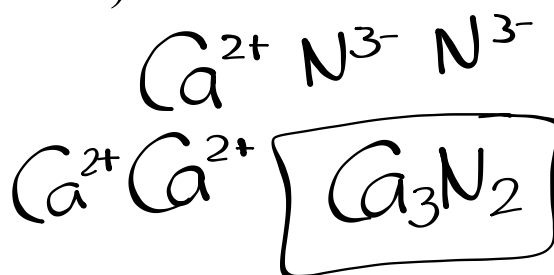


Write the chemical formula or name for the following ionic compounds:

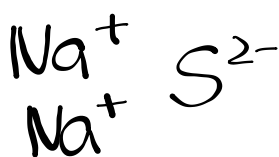
a) magnesium oxide



b) calcium nitride



c) Na_2S



Sodium sulfide

d) CaF_2

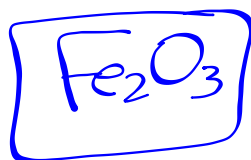
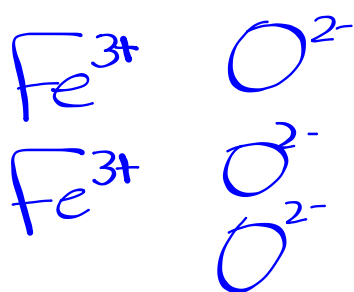


Calcium fluoride

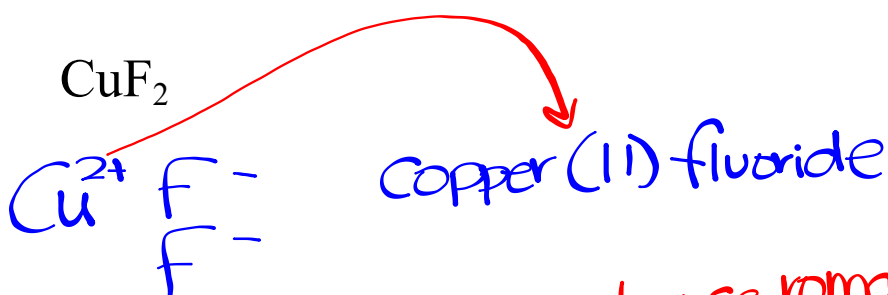
Type II Binary Ionic Compounds

Ex. Name to Formula:

iron (III) oxide



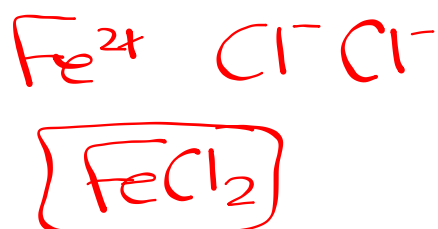
Formula to Name :



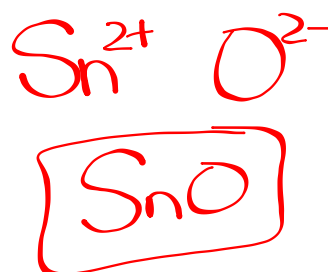
* only use roman numeral
if metal is multi-
valent.

Write the chemical formula or name for the following ionic compounds:

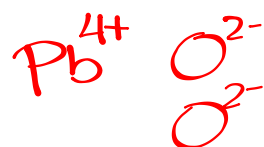
a) iron (II) chloride



b) tin (II) oxide

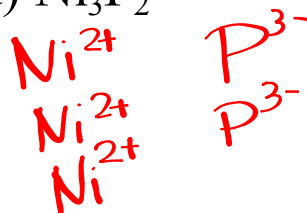


c) PbO_2



lead(IV) oxide

d) Ni_3P_2



nickel(II) phosphide

Binary Ionic Compounds Type II

Worksheet