

Check Homework - Worksheet

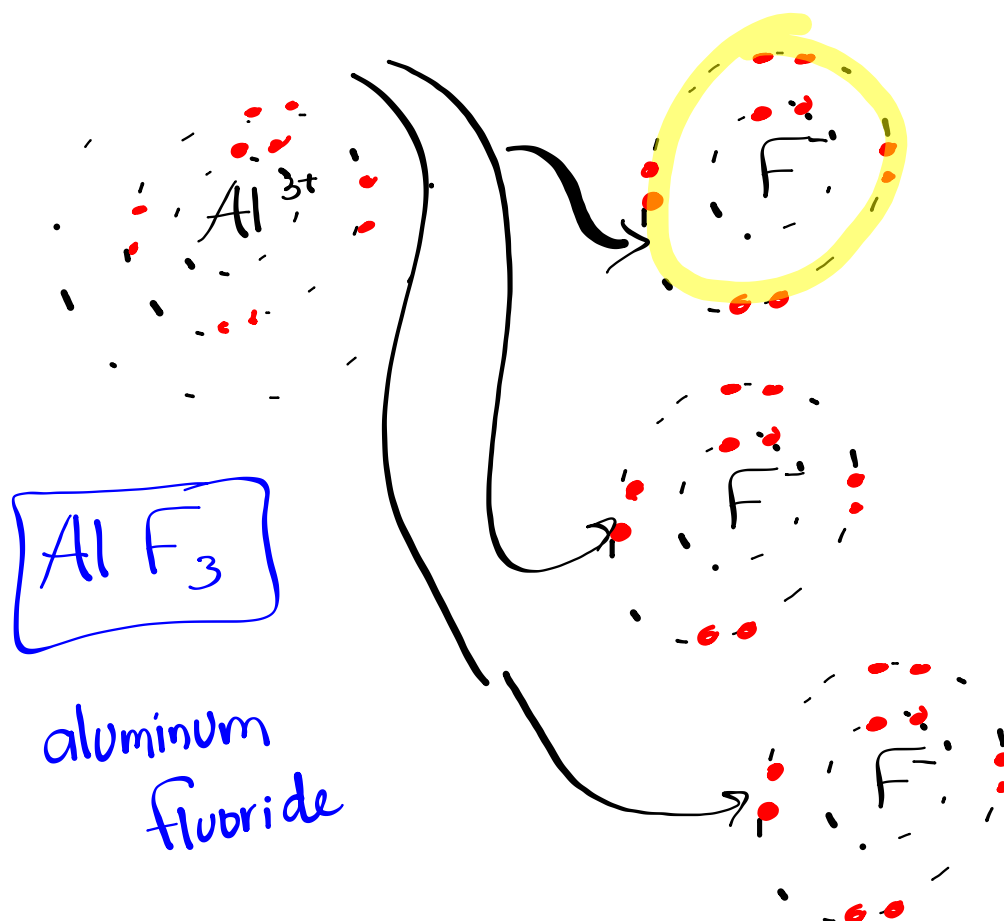
Worksheet - Ionic Compounds

IONIC COMPOUNDS

Ionic compounds are made by elements transferring or sharing electrons.

- the further an e⁻ is away from the nucleus, the greater the possibility of it making a compound with another element
- the **outermost electrons** are involved in making compounds

Ex. aluminum and fluorine



Ionic Compounds

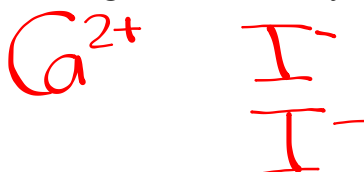
How to Write the Formula for an Ionic Compound

Ex. Calcium and iodine

1. Write the symbols, with the metal always being written first.



2. Write the ionic charge above the symbol to indicate the stable ion that each element forms



3. Determine how many ions of each type you need so that the total ionic charge is zero.

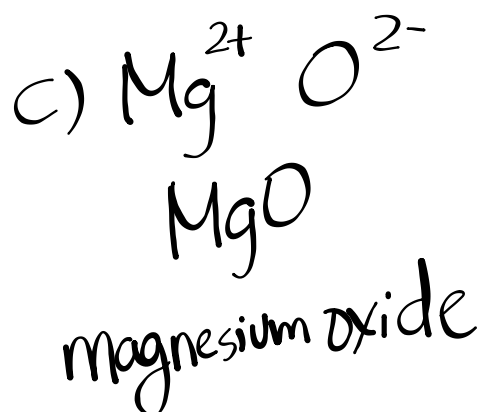
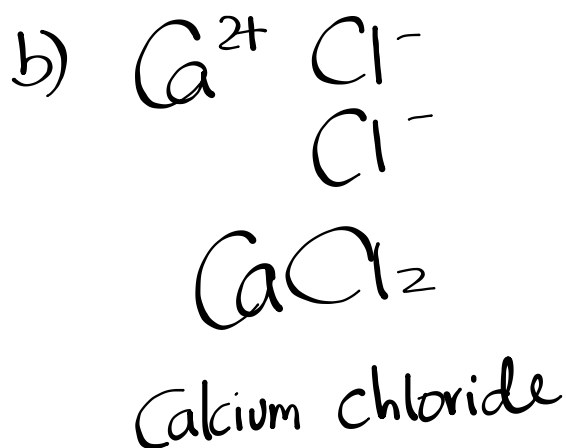
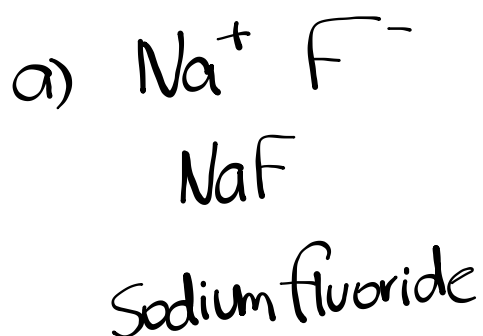


4. Write the formula using subscripts to indicate the number of ions of each type.

- when naming ionic compounds, the name of the metal remains the same but the name of the nonmetal changes to an -ide ending

Write the formulas and the names of the compounds formed by the combination of the elements:

- a) sodium and fluorine
- b) calcium and chlorine
- c) magnesium and oxygen



Nowwork

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