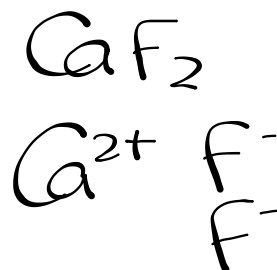
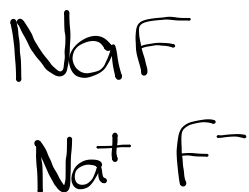
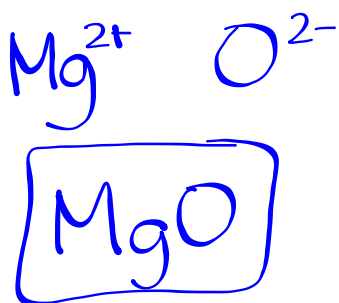


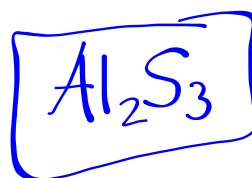
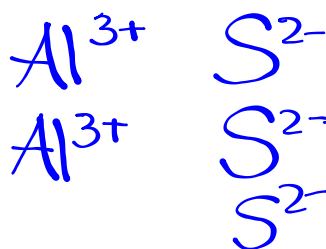
## Check Homework - #1-6



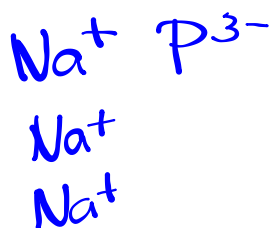
5. b) magnesium oxide



d) aluminum sulfide

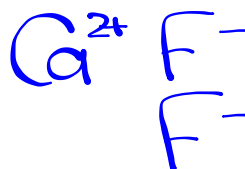


6. b)  $\text{Na}_3\text{P}$



Sodium phosphide

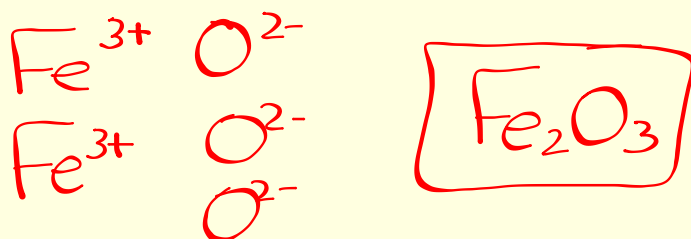
c)  $\text{CaF}_2$



Calcium fluoride

## Multivalent Metals

- some metals have more than one charge they are called **multi-valent ions**
  - these elements are found in the middle block of the periodic table
  - the charge that is to be used is indicated in brackets with a Roman numeral (Table 2 p. 195)
- Ex. iron(III) oxide



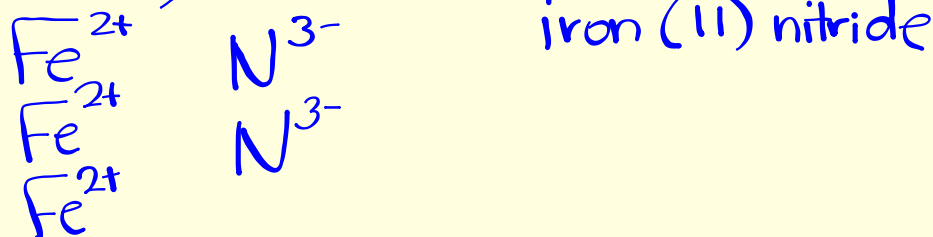
*Naming Ionic Compounds from Formula (multivalent ions):*

- Identify positive ion (metal) and negative ion (nonmetal)
- If metal is multivalent, determine its charge from the formula (balance total positives and negatives) and include in name

Ex. CuCl

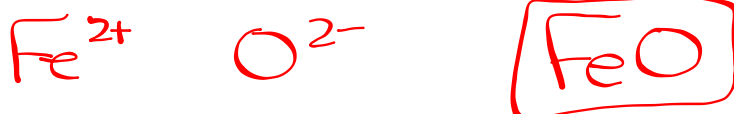


Ex.  $\text{Fe}_3\text{N}_2$

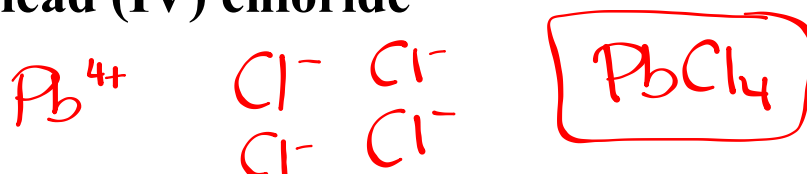


Write the chemical formula of the following ionic compounds:

(a) iron (II) oxide



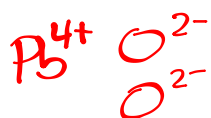
(b) lead (IV) chloride



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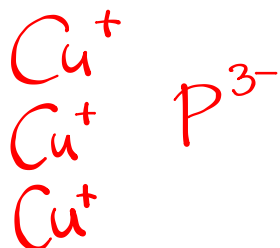
Write the name of the following ionic compounds:

(a)  $\text{PbO}_2$



lead (IV) oxide

(b)  $\text{Cu}_3\text{P}$



Copper (I) phosphide

# Homework

p.195 #7-10

Ionic Compounds Sheet #1