

Ionic and Molecular Compounds Review Worksheet

Write the name for each of the following Ionic Compounds:

1. NaBr _____

2. $\text{Ti}(\text{SO}_4)_2$ _____

3. FePO_4 _____

4. K_3N _____

5. CuOH _____

6. VCl_4 _____

7. Mg_3P_2 _____

8. $\text{Sr}(\text{HCO}_3)_2$ _____

9. Ba_3N_2 _____

10. $\text{Zn}(\text{NO}_3)_2$ _____

Write the name for each of the following Molecular Compounds:

1. CO _____

2. P_2O_5 _____

3. CO_2 _____

4. CH_4 _____

5. CF_4 _____

6. SO_2 _____

7. N_2O _____

8. N_2O_4 _____

9. BCl_3 _____

10. C_2S_4 _____

Write the formula for each of the following Ionic Compounds:

1. nickel (III) sulfide _____

2. copper (II) sulfate _____

3. manganese (II) phosphate _____

4. aluminum phosphate _____

5. magnesium hydroxide _____

6. potassium carbonate _____

7. silver oxide _____

8. tin (IV) selenide _____

9. lithium nitride _____

10. copper (I) sulfide _____

Write the formula for each of the following Molecular Compounds:

1. silicon dioxide _____

2. diboron tetrabromide _____

3. carbon tetrachloride _____

4. nitrogen monoxide _____

5. sulfur dioxide _____

6. hydrogen peroxide _____

7. dicarbon tetrasulfide _____

8. nitrogen tribromide _____

9. methane _____

10. oxygen _____