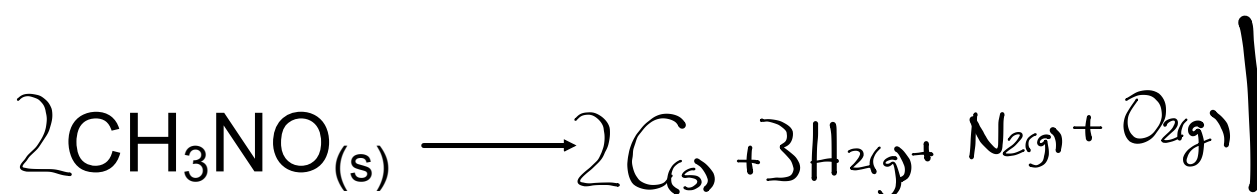
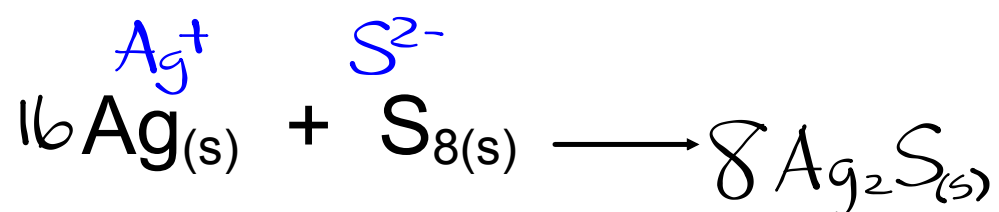


Warm Up



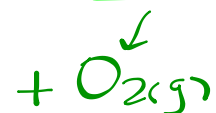
Check Homework - Worksheet

Chemical Reactions

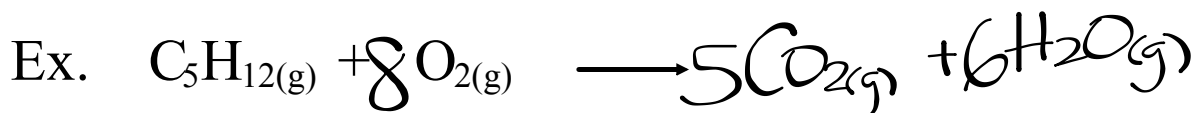
III. Combustion Reaction

A complete combustion reaction is the burning of a **substance with oxygen** to produce the most common oxides of the elements in the substance being **burned**.

Most Common Oxides:

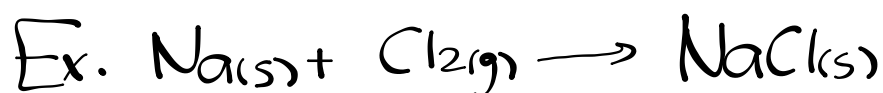


- Carbon : $\text{CO}_{2(g)}$
- Hydrogen: $\text{H}_2\text{O}_{(g)}$
- Sulfur: $\text{SO}_{2(g)}$
- Nitrogen: $\text{NO}_{2(g)}$
- A metal: Oxide of metal with most common ion charge



I. FORMATION (COMBINATION)

element + element \rightarrow compound

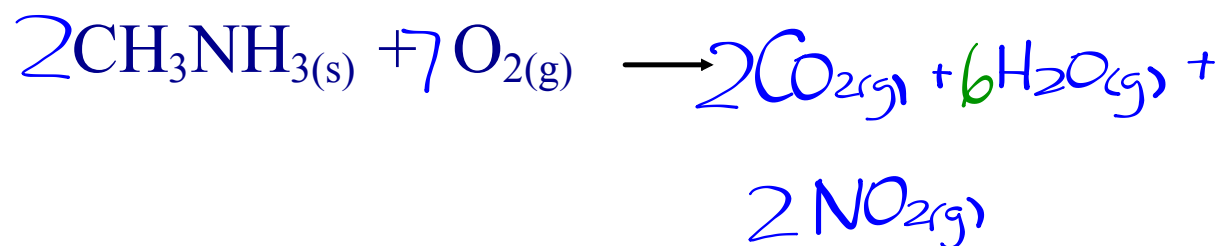
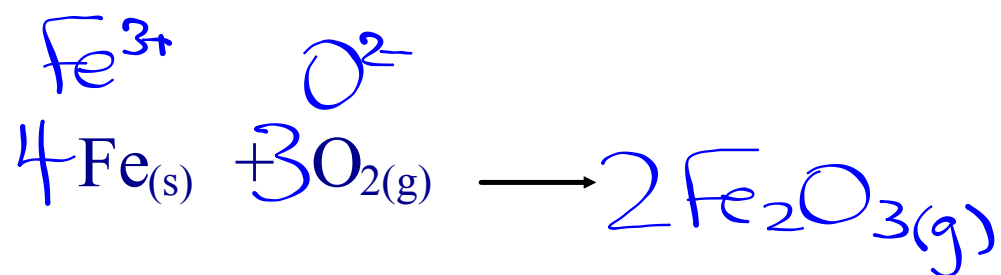


II. DECOMPOSITION

compound \rightarrow elements

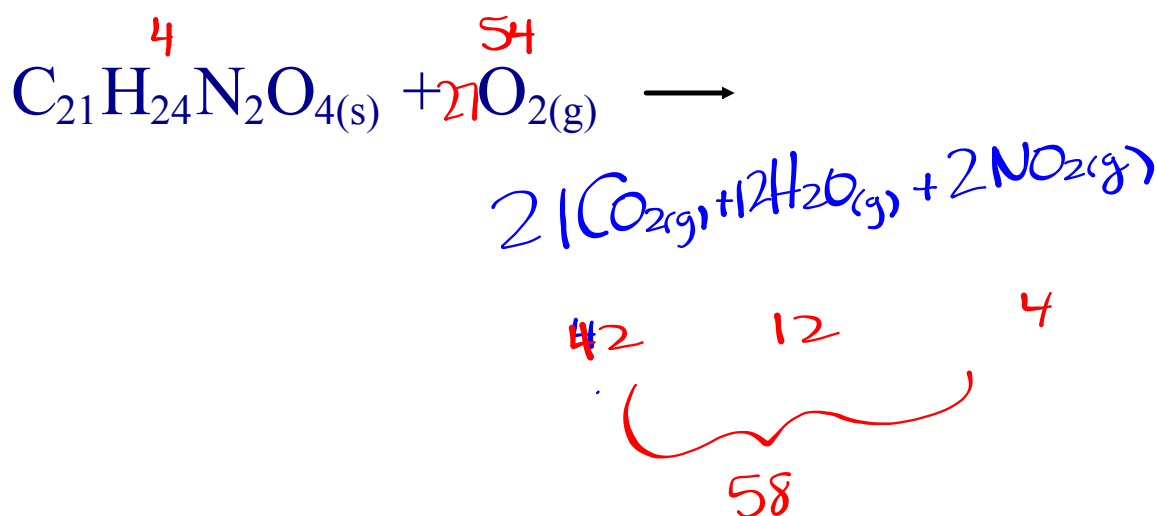
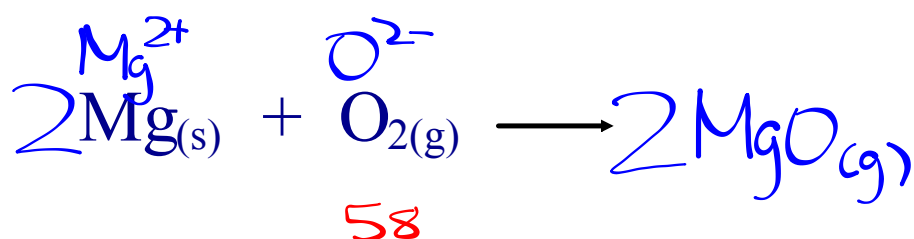
III. COMBUSTION

element / compound + $\text{O}_2(\text{g}) \rightarrow$ most common oxides



Combustion Reactions

Write a balanced chemical equation for the following combustion reactions:



Combustion of C_4H_{10}



Homework

p. 331 #13, 14

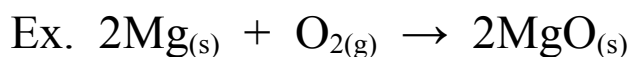
p. 332 #15, 16

p. 337 #20, 21

Chemical Reactions

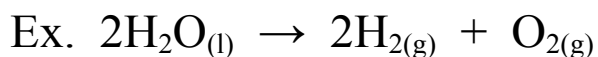
I. Formation Reactions

elements compound



II. Decomposition Reactions

compound elements



III. Combustion Reaction

substance + oxygen \longrightarrow most common oxides

