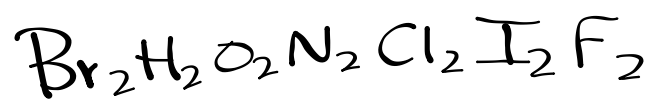
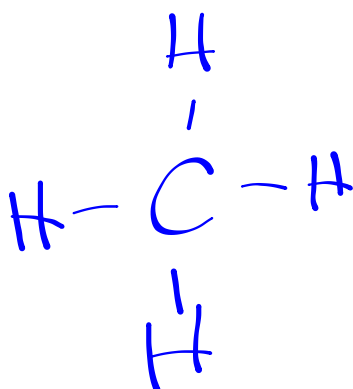


## Unit 2 - Compounds

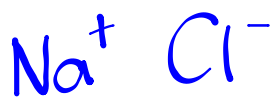
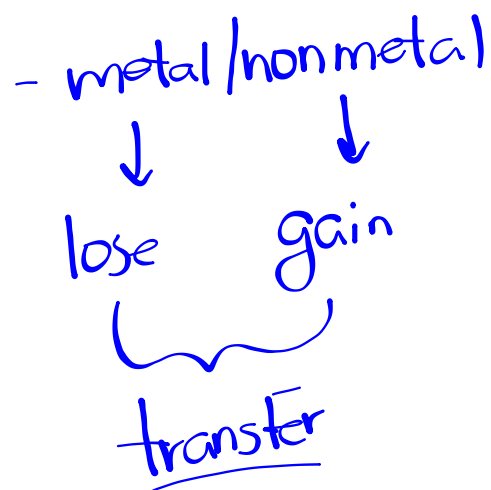
- Properties of Ionic Compounds, Molecular Compounds, Acids, and Bases (Empirical and Theoretical)
- Naming Ionic Compounds
- Writing formulas for Ionic Compounds
- Ionic hydrates
- Naming Molecular Compounds
- Writing formulas for Molecular Compounds
- Molecular Elements *Br HONClIF / P<sub>4</sub> S<sub>8</sub>*
- Drawing structural diagrams
- Naming and writing formulas for Acids and Bases
- Lab - Identifying Unknown Compounds

~~Q~~

CH<sub>4</sub> → molecular (nonmetals)

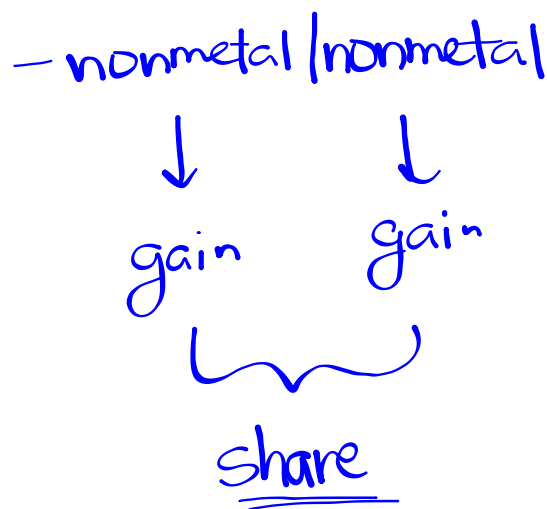


# IONIC



+/- attraction

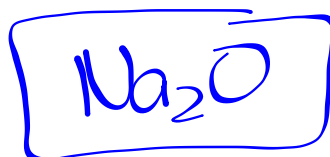
# COVALENT

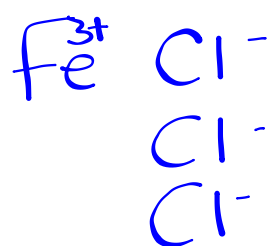




Calcium chloride

sodium oxide





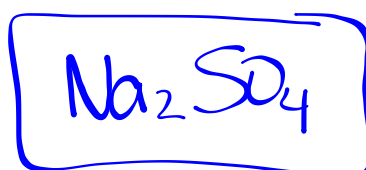
iron(III) chloride

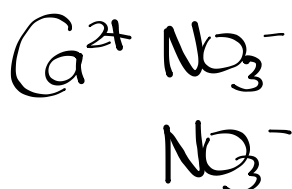
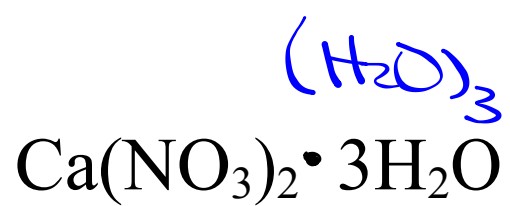
copper (II) sulfide



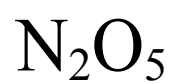
lithium nitrate

sodium sulfate





Calcium nitrate trihydrate  
" - 3-water



molecular

dinitrogen pentoxide



HClO

$H^+$   $ClO^-$  hypochlorous acid  
aqueous hydrogen hypochlorite  
or

-ate  
chromic acid

$H^+$   $CrO_4^{2-}$   
 $H^+$



**Review Questions p. 281-282**

**Worksheets**

**# 43-61, 65-71**