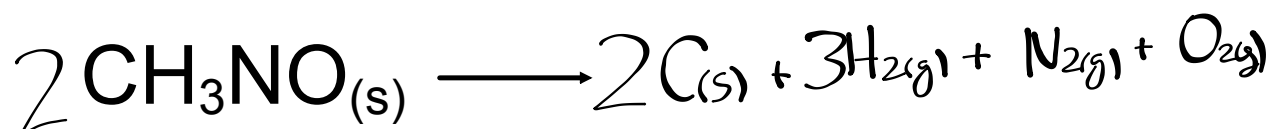
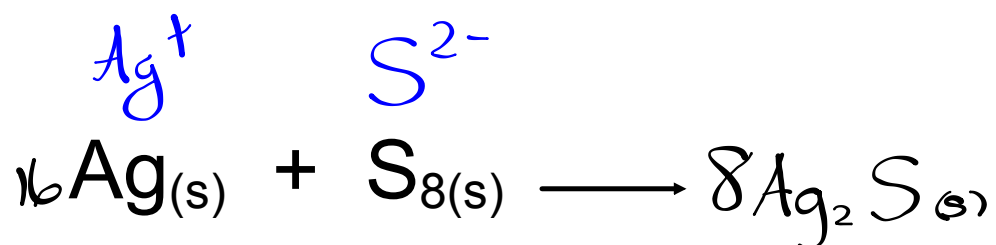
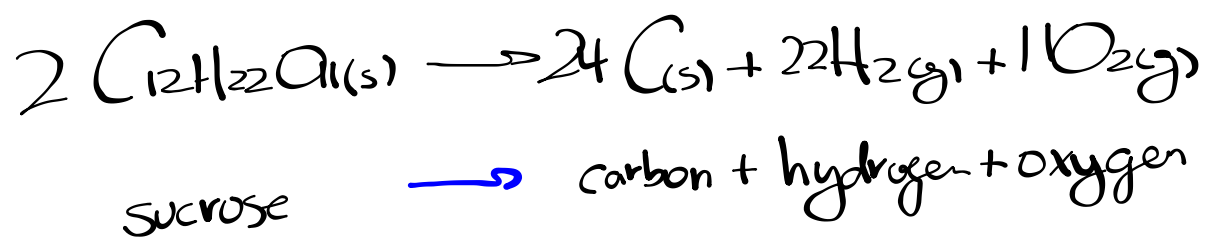


Warm Up



Check Homework - Worksheet

DECOMP.



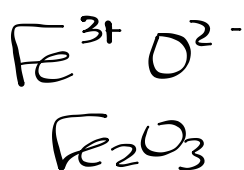
Chemical Reactions

III. Combustion Reaction

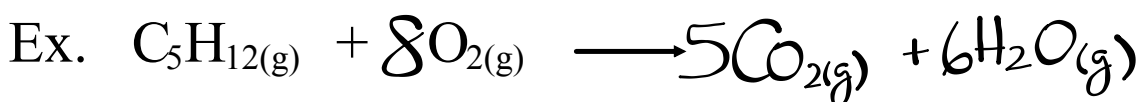
A complete combustion reaction is the **burning** of a substance **with oxygen** to produce the most common oxides of the elements in the substance being burned.

Most Common Oxides:

- Carbon : $\text{CO}_{2(g)}$
- Hydrogen: $\text{H}_2\text{O}_{(g)}$
- Sulfur: $\text{SO}_{2(g)}$
- Nitrogen: $\text{NO}_{2(g)}$
- A metal: Oxide of metal with most common ion charge



* products are gases !



I. FORMATION

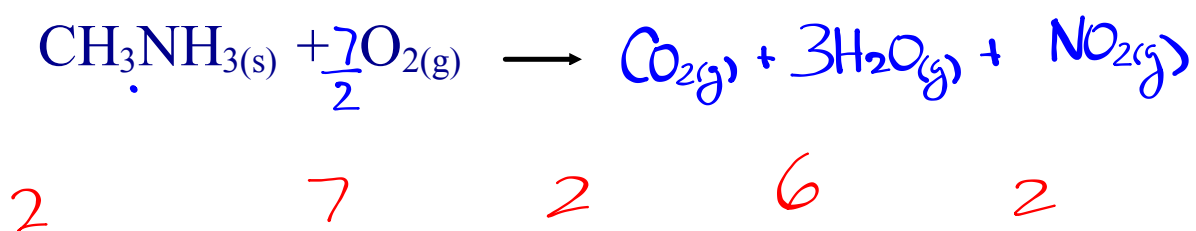
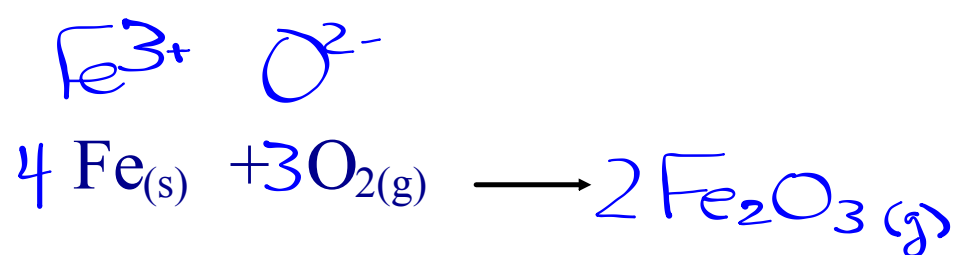
element + element \rightarrow compound

II. DECOMPOSITION

Compound \rightarrow elements

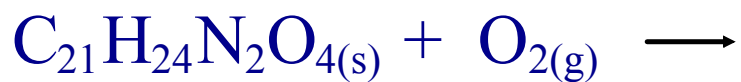
III. COMBUSTION

element/compound + oxygen(O_2) \rightarrow most common oxides

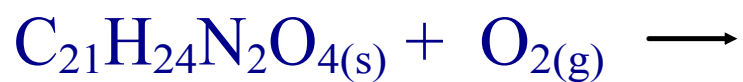


Combustion Reactions

Write a balanced chemical equation for the following combustion reactions:



Homework



p. 331 #13, 14

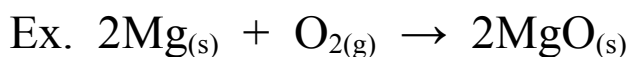
p. 332 #15, 16

p. 337 #20, 21

Chemical Reactions

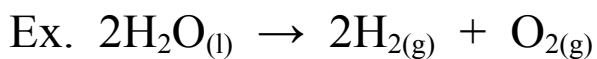
I. Formation Reactions

elements compound



II. Decomposition Reactions

compound elements



III. Combustion Reaction

substance + oxygen \longrightarrow most common oxides

