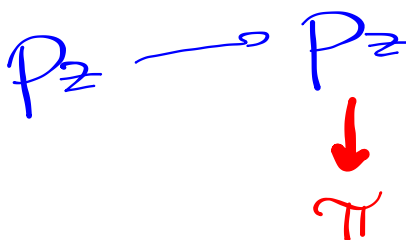
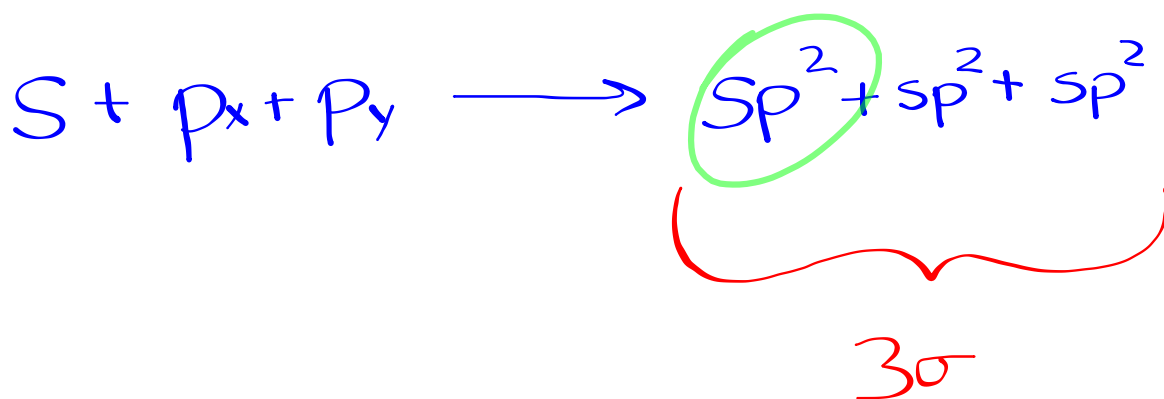
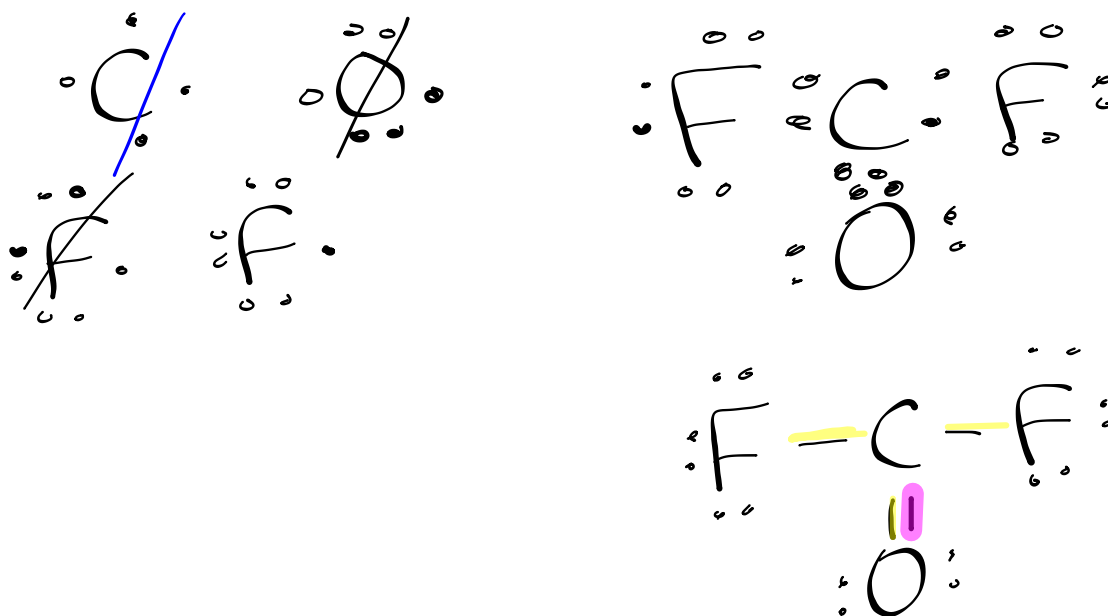


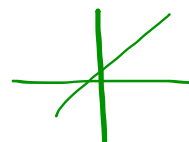
Warm Up

Determine the type of hybrid orbitals used by the carbon atom in CF_2O . How many sigma and pi bonds does carbon make?



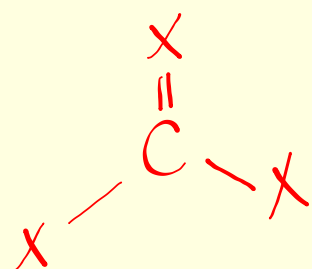
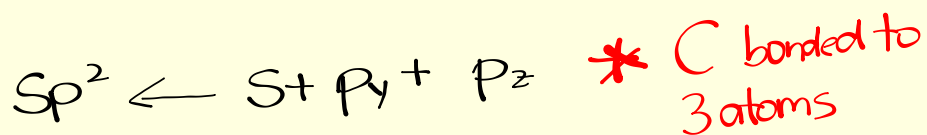
2p 1 1 1
2s 1
1s 1

P_x P_y P_z

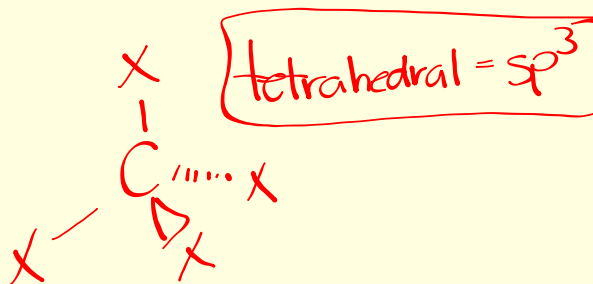
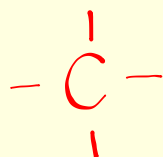
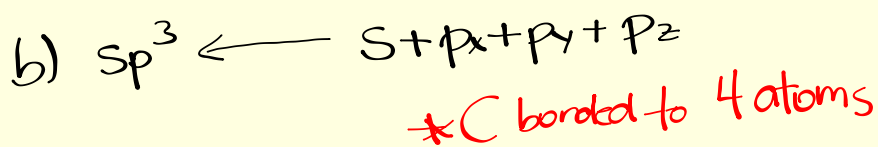


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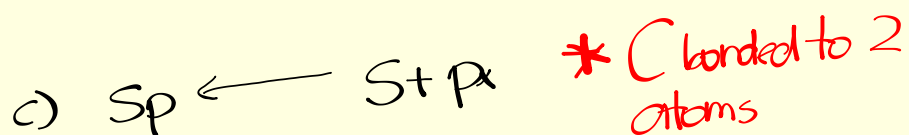
26.a) Shape? tetrahedral, trigonal planar
linear



trigonal planar = sp^2



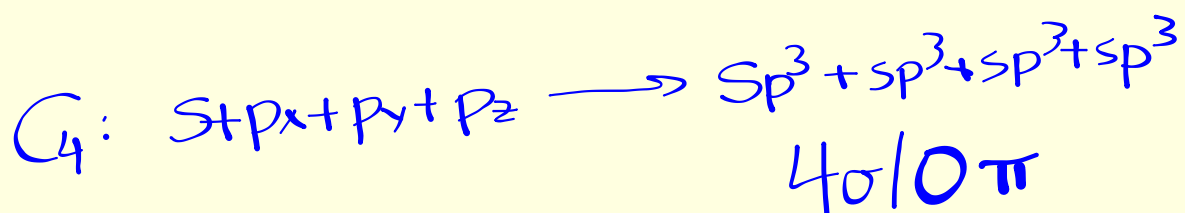
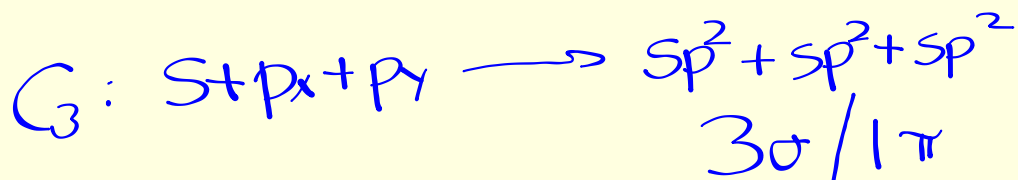
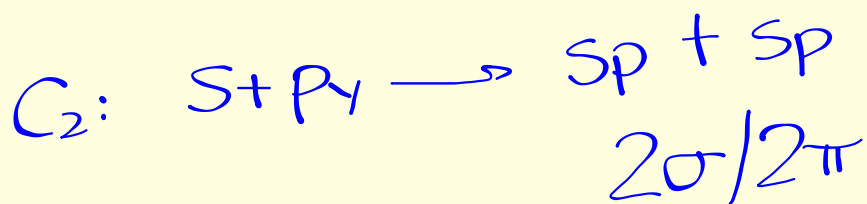
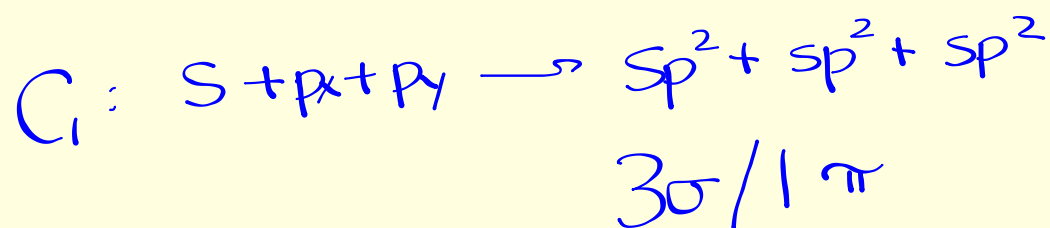
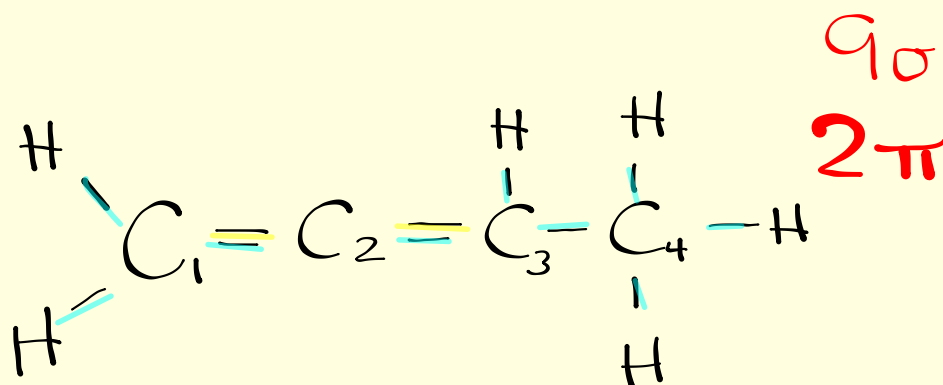
tetrahedral = sp^3



linear = sp



- 1) Determine the type of hybrid orbitals used by each carbon atom.
- 2) State the number of sigma and pi bonds used by each C atom.
- 3) State the total number of sigma and pi bonds.



Worksheet 8.3