


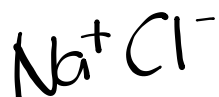
Unit 2 - Compounds

- Properties of Ionic Compounds, Molecular Compounds, Acids, and Bases (Empirical and Theoretical)
- Naming Ionic Compounds
- Writing formulas for Ionic Compounds
- Ionic hydrates
- Naming Molecular Compounds
- Writing formulas for Molecular Compounds
- Molecular Elements **Br Hg N Cl I F (P₄, S₈)**
- Drawing structural diagrams
- Naming and writing formulas for Acids and Bases
- Lab - Identifying Unknown Compounds

IONIC




transfer



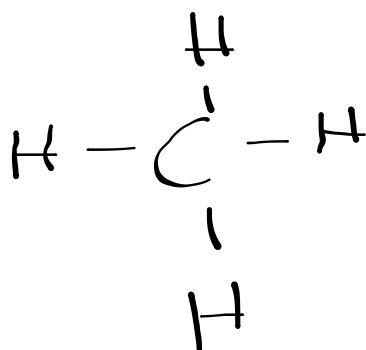
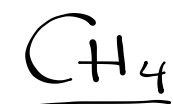

+/- attraction

COVALENT

nonmetals



Share



Review Questions p. 281-282

Worksheets

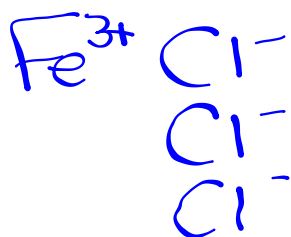
43-61, 65-71



Calcium Chloride

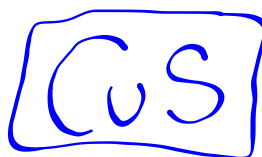
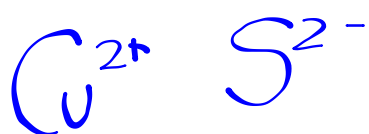
sodium oxide





iron (III) chloride

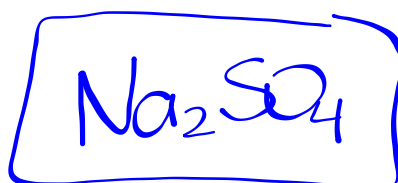
copper (II) sulfide

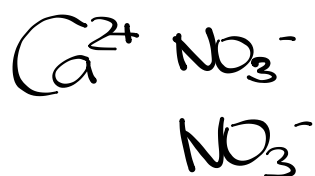
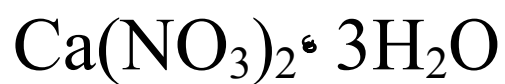




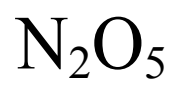
lithium nitrate

sodium sulfate





Calcium nitrate trihydrate
-3-water



dinitrogen pentoxide

HClO

$H^+ ClO^-$

aqueous hydrogen hypochlorite

hypochlorous acid

chromic acid

$H^+ CrO_4^{2-}$

H^+

H_2CrO_4