Homework

$\Delta H = -q$

② $m = 100.g$
$C = ?$
$T_i = 100.°C$
H_2O
$m = 100.0g$
$T_i = 20.0°C$
$T_f = 23.395°C$ *

$q_4 = -q_{H_2O}$

$mC\Delta T = -mC\Delta T$

$(100g)C(-76.605°C) =$
$-(100.0g)(4.19\frac{J}{g\cdot°C})(3.395°C)$

$C = \frac{-(100.0g)(4.19\frac{J}{g\cdot°C})(3.395°C)}{(100.g)(-76.605°C)}$

$C = 0.186\frac{J}{g\cdot°C}$

Homework

Worksheet