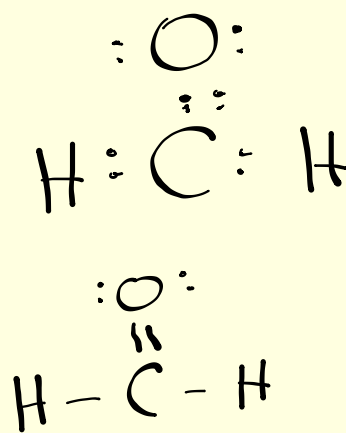
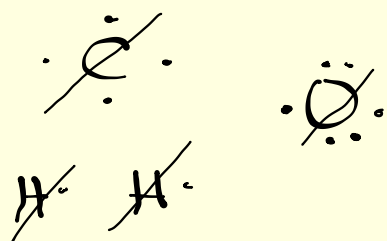


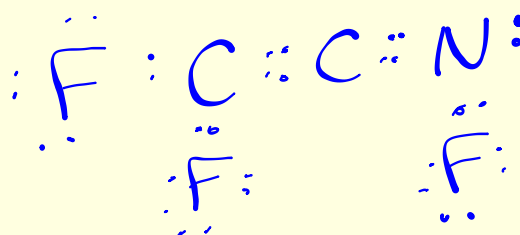
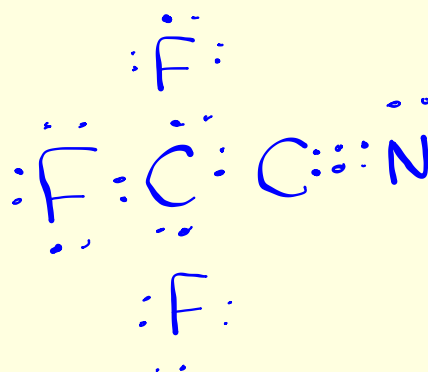
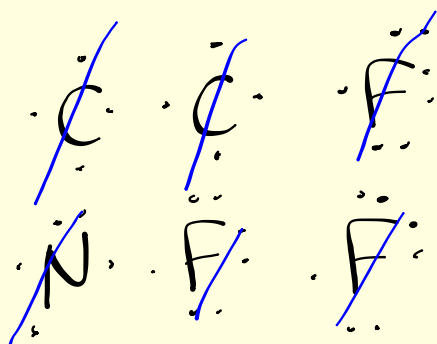
## Warm Up

Draw an electron dot structure and structural diagram for the following:

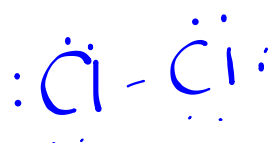
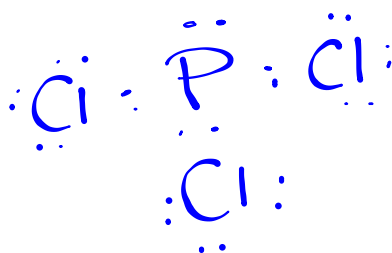
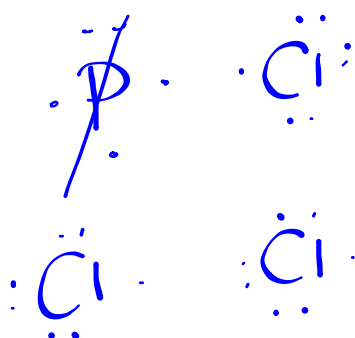
a)  $\text{CH}_2\text{O}$



b)  $\text{C}_2\text{NF}_3$



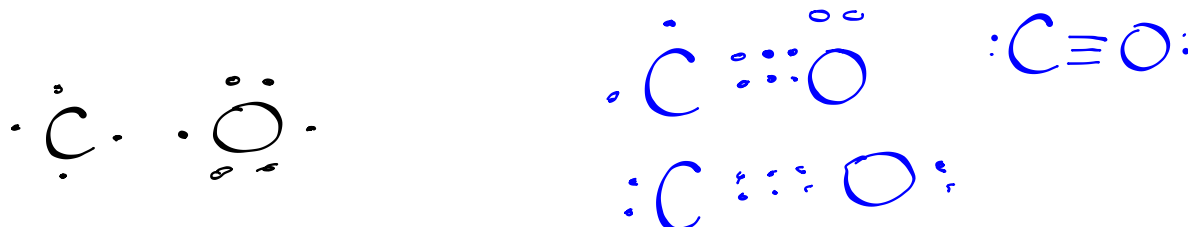
## p. 225 #7,8

7.a)  $\text{Cl}_2$ 8.b)  $\text{PCl}_3$ 

|

## Coordinate Covalent Bonds

CO

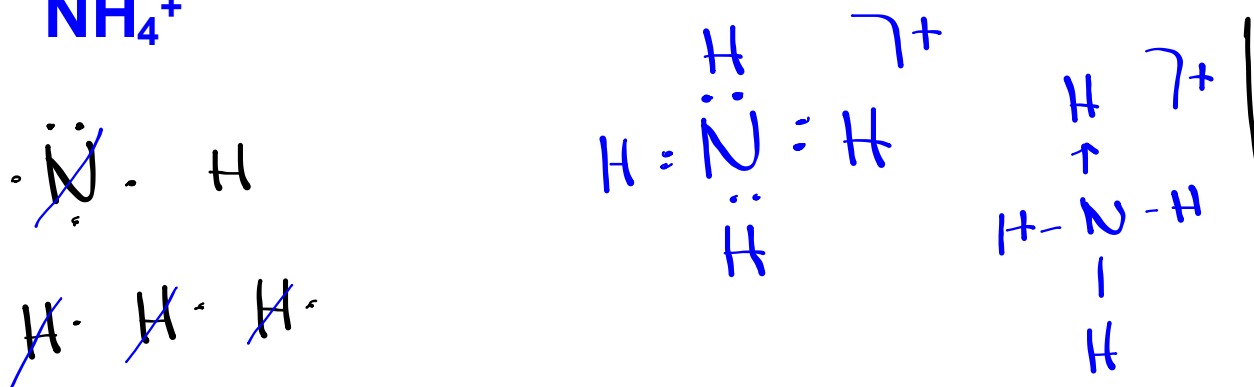


### Coordinate Covalent Bond

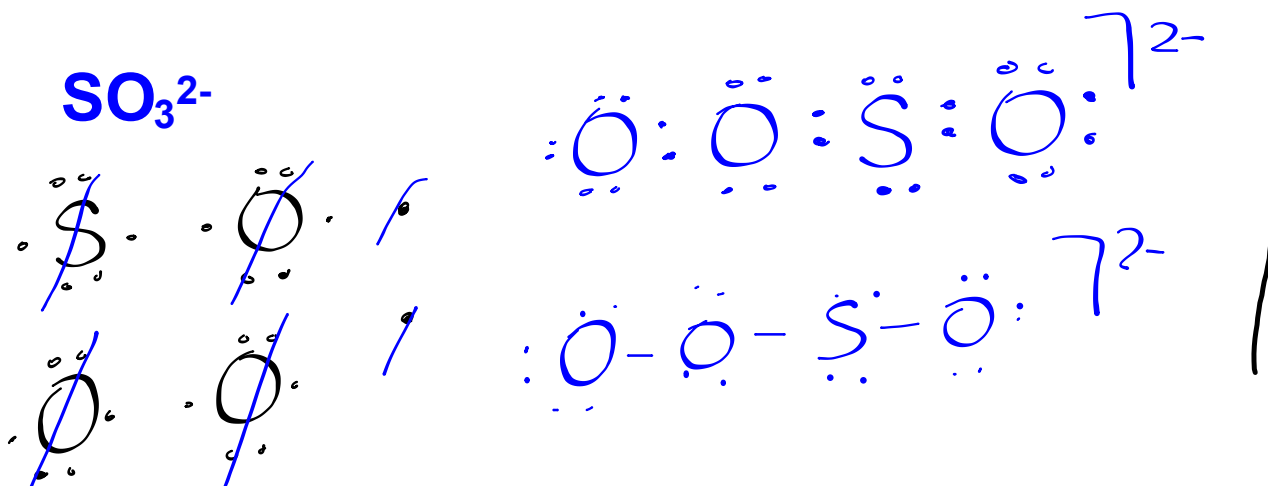
A covalent bond in which one atom contributes a shared pair of electrons.

- Coordinate covalent bonding is common in polyatomic ions.

NH<sub>4</sub><sup>+</sup>



SO<sub>3</sub><sup>2-</sup>



# Homework

**p. 225 #9-12**  
**Worksheet**