

Electron Configurations

Quantum Mechanical Model of an Atom

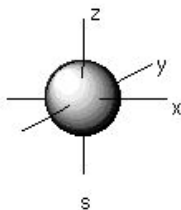
The quantum mechanical model determines the allowed energies an electron can have and how likely it is to find the electron in various locations around the nucleus.

atomic orbital - region of space in which there is a high probability to find an electron

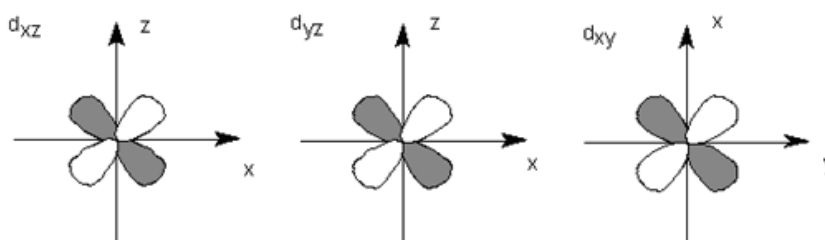
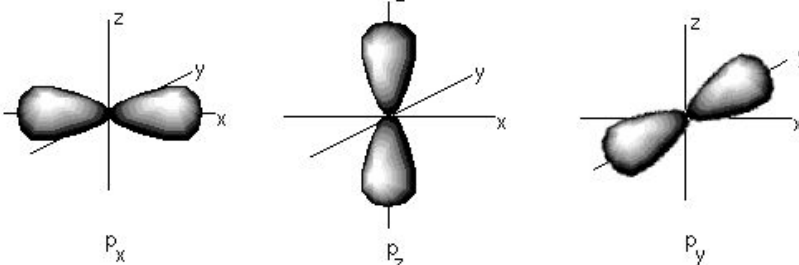
Principal quantum numbers (n) represent energy levels of electrons (i.e., $n = 1, 2, 3, 4$, etc.)

There may be several orbitals with different shapes at different energy levels.

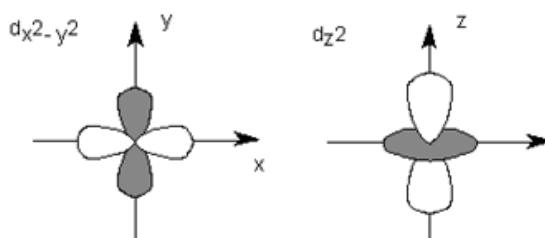
s orbital



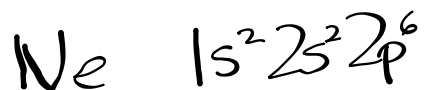
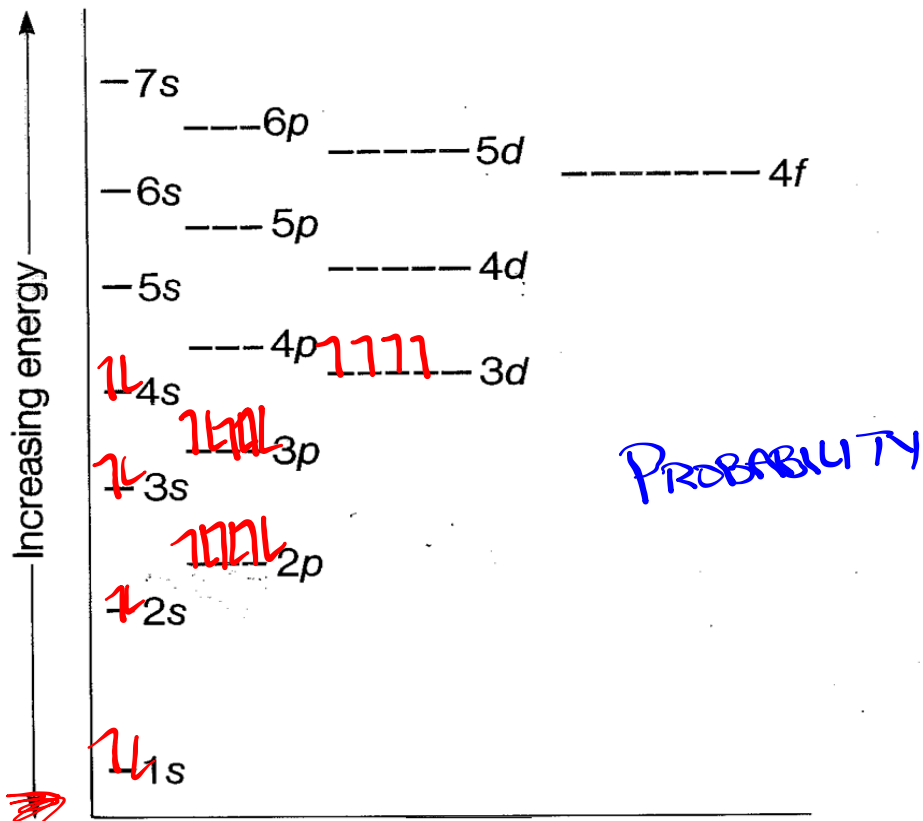
p orbitals



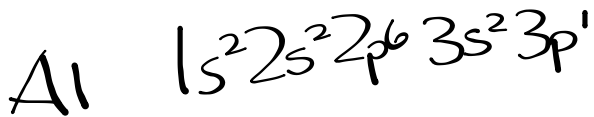
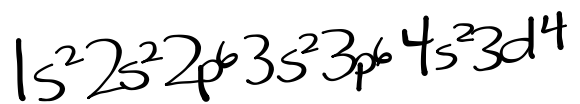
d orbitals



Aufbau Diagram

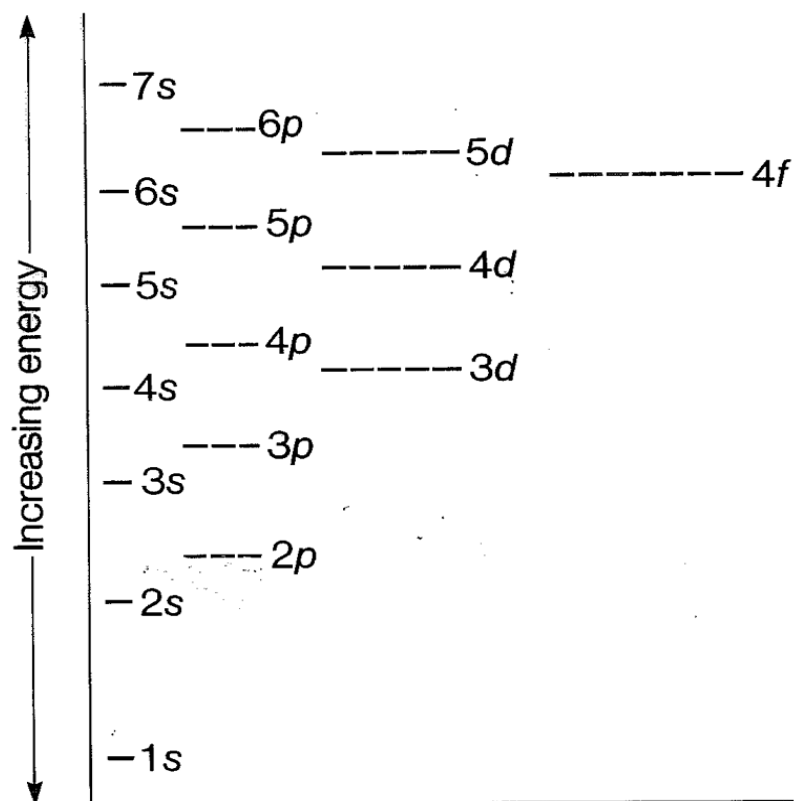


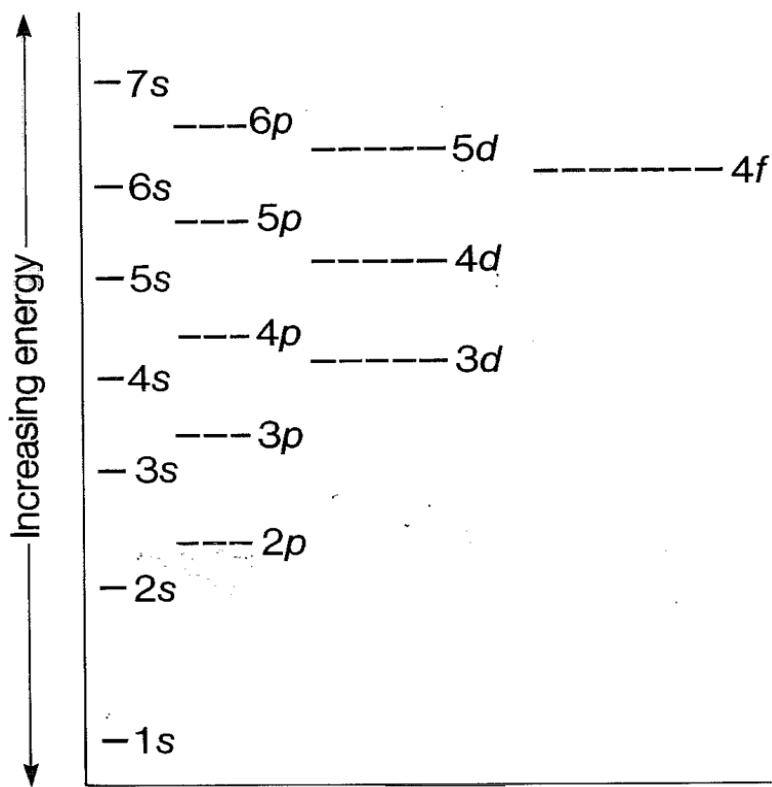
(10)



(13)



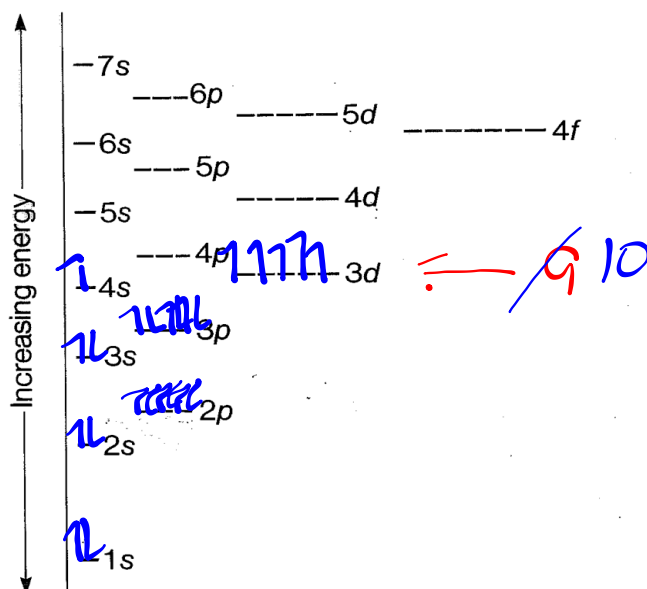




Exceptional Electron Configurations

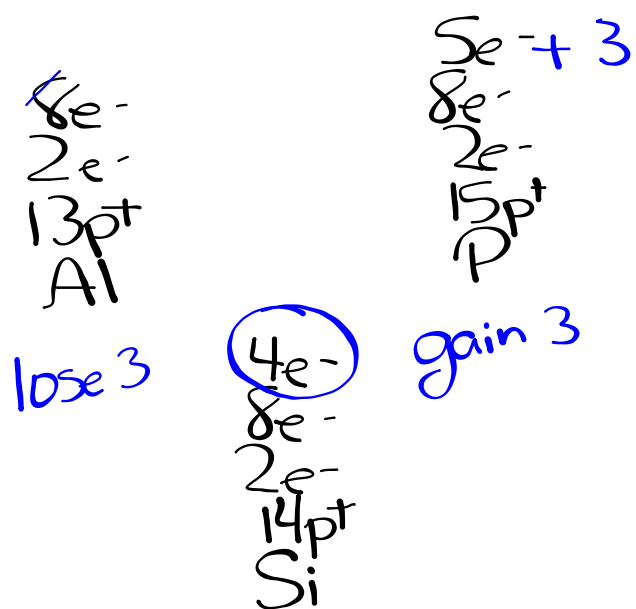
Although half-filled sublevels are not as stable as filled sublevels, they are more stable than other configurations.

- more likely to occur at higher principal quantum numbers, because of the small energy differences between sublevels.



Cu
(29)

Cr $1s^2 2s^2 2p^6 3s^2 3p^6 3d^4 4s^2$
(24)




Homework

Section 5.1 & 5.2 p. 128 - 137

p. 131 #1-7

p. 135 #8-9

p. 136 #10-13

 <http://www.chalkbored.com/lessons/chemistry-12/periodic-configurations.pdf>