## Warm Up

$$3 \text{ Fe}_{(s)} + \text{Al}_2(SO_4)_{3(aq)} \longrightarrow 2 \text{Al}_{(s)} + 3 \text{FeSO}_{4(aq)}$$

## **Check Homework - Worksheet**

## **Types of Chemical Reactions**

There are five types of chemical reactions:

- I. Formation/Combinationreactions occur when two substances (normally elements) react to form an ionic or molecular compound
- -when a metal and nonmetal react, the product will be the ionic compound formed by the most common ions.

$$\begin{array}{c} \text{Mg}^{2+} & \bigcirc^{2-} \\ \text{Ex2Mg}_{(s)} + \text{O}_{2(g)} & \rightarrow 2 \text{MgO}_{(s)} \\ \text{elements} & \text{compound} \\ \text{reactants} & \text{product} \end{array}$$

II. A decomposition reaction is the result of an ionic or molecular compound breaking down into its elements. ⇒it is the reverse of a formation reaction

Ex2H<sub>2</sub>O<sub>(l)</sub> 
$$\rightarrow$$
2H<sub>2</sub>G + O<sub>2G</sub> compound reactant elements products

$$Pb^{2t}$$
  $O^{2t}$   $O^{2t}$ 

Cu<sup>24</sup> 02b) copper (11) oxide -> CuOGI ->

## Homework

Worksheet

p. 331 #13, 14

p. 332 #15, 16