April 30, 2019

answers pg 198 #4,6,7/ Polyatomic WS
review ALL Ionic Compounds

Quiz Tomorrow!
Test Tuesday May 7 !!

Warm-Up

Name the following ionic compounds:

LiOH lithium hydroxide

$Ba(NO_3)_2$ barium nitrate

$CaSO_4$ calcium sulfate

Give the formula for following ionic compounds:

beryllium hydroxide $Be^{2+}OH^- \quad Be(OH)_2$

magnesium chlorate $Mg^{2+}ClO_3^- \quad Mg(ClO_3)_2$

aluminum phosphate $Al^{3+}PO_4^{3-} \quad AlPO_4$

4. (a) potassium carbonate
(b) sodium sulfate
(c) aluminum bicarbonate
(d) silver nitrate

6. A polyatomic ion chemically acts as a single particle. Therefore, the presence of a polyatomic ion must be clearly indicated in the formula of the compound. Both ammonium (NH_4^+) and nitrate (NO_3^-) are polyatomic ions.

7. (a) ammonium chloride, NH_4Cl
(b) ammonium sulfate, $(NH_4)_2SO_4$

Polyatomic Compounds: Names and Formulas Worksheet

(a) magnesium sulfate $MgSO_4$

(b) sodium chlorate $NaClO_3$

(c) aluminum nitrate $Al(NO_3)_3$

(d) potassium hydroxide KOH

(e) lithium phosphate Li_3PO_4

(f) calcium carbonate $CaCO_3$

(g) beryllium sulfate $BeSO_4$

(h) sodium bicarbonate $NaHCO_3$

(i) magnesium hydroxide $Mg(OH)_2$

(j) aluminum phosphate $AlPO_4$

Oct 30-8:43 AM

(k) copper(I) chlorate $CuClO_3$

(l) calcium sulfate $CaSO_4$

(m) nitric acid HNO_3

(n) carbonic acid H_2CO_3

(o) sulfuric acid H_2SO_4

(p) lead(II) nitrate $Pb(NO_3)_2$

(q) phosphoric acid H_3PO_4

(r) copper(II) hydroxide $Cu(OH)_2$

(s) iron(II) phosphate $Fe_3(PO_4)_2$

(t) calcium chlorate $Ca(ClO_3)_2$

2. Write the names for the following compounds

2. Write the names for the following compounds.

(a) Li_2CO_3 lithium carbonate

(b) $AlHCO_3$ aluminum bicarbonate

(c) $Mg_3(PO4)_2$ magnesium phosphate

(d) $Ca(NO_3)_2$ calcium nitrate

(e) K_2SO_4 potassium sulfate

(f) $HNO_{3(aq)}$ nitric acid

(g) $NaNO_3$ sodium nitrate

(h) $Al(OH)_3$ aluminum hydroxide

(i) $CuSO_4$ copper(II) sulfate

(j) $Fe(ClO)_3$ iron (III) hypoclorite

1+ 2+ 2- 2+ 3+ 1-

2+ 3+ 3-

$K)Pb_3(PO_4)_2$ lead (II) phosphate

(l) $Sn(ClO_3)_2$ tin (II) chlorate

(m) $NaOH$ sodium hydroxide

(n) $H_3PO_{4(aq)}$ phosphoric acid

(o) $H_2CO_{3(aq)}$ carbonic acid

(p) $CuNO_3$ copper (I) nitrate

(q) $H_2SO_{4(aq)}$ sulfuric acid

r) $FeSO_4$ iron (II) sulfate

s) $Ca(HCO_3)_2$ calcium bicarbonate hydrogen carbonate

t) K_3PO_4 potassium phosphate

Review Ionic Compounds

Regular

Write Name:		Write Formula:	
LiCl	lithium chloride	sodium sulfide	Na_2S
Mg_3S_2	magnesium sulfide	potassium bromide	KBr

Multivalent

FeO	iron(II) oxide	maganese (IV) oxide	MnO_2
$CuCl_2$	copper(II) chloride	vanadium (V) nitride	V_3N_5

Polyatomic

$Mg_2(SO_4)_3$	magnesium sulfate	calcium phosphate	$Ca_3(PO_4)_2$
$Ca(OH)_2$	calcium hydroxide	beryllium sulfate	$BeSO_4$

Dec 6-1:22 PM

nitride N^{3-}

nitrate NO_3^{1-}

Review WS ALL Ionic Compounds

Dec 6-1:31 PM