

April 30, 2019

answers pg 198 #4,6,7/ Polyatomic WS
review ALL Ionic Compounds

Quiz Tomorrow!

Test Tuesday May 7 !!

Warm-Up

Name the following ionic compounds:



lithium hydroxide



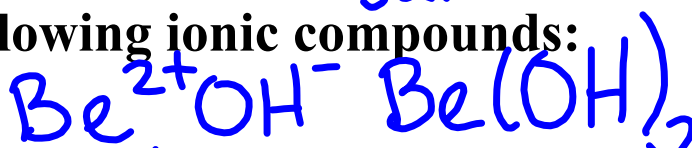
barium nitrate



calcium sulfate

Give the formula for following ionic compounds:

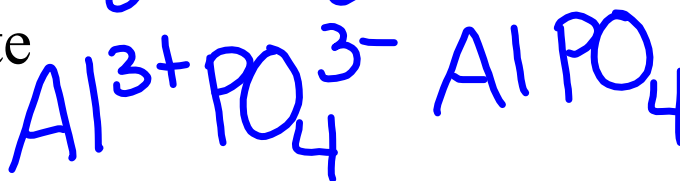
beryllium hydroxide



magnesium chlorate



aluminum phosphate



4. (a) potassium carbonate
- (b) sodium sulfate
- (c) aluminum bicarbonate
- (d) silver nitrate

6. A polyatomic ion chemically acts as a single particle. Therefore, the presence of a polyatomic ion must be clearly indicated in the formula of the compound. Both ammonium (NH_4^+) and nitrate (NO_3^-) are polyatomic ions.

7. (a) ammonium chloride, NH_4Cl
- (b) ammonium sulfate, $(\text{NH}_4)_2\text{SO}_4$

Polyatomic Compounds: Names and Formulas Worksheet

- (a) magnesium sulfate MgSO₄
- (b) sodium chlorate NaClO₃
- (c) aluminum nitrate Al(NO₃)₃
- (d) potassium hydroxide KOH
- (e) lithium phosphate Li₃PO₄
- (f) calcium carbonate CaCO₃
- (g) beryllium sulfate BeSO₄
- (h) sodium bicarbonate NaHCO₃
- (i) magnesium hydroxide Mg(OH)₂
- (j) aluminum phosphate AlPO₄

(k) copper(I) chlorate	<u>$CuClO_3$</u>
(l) calcium sulfate	<u>$CaSO_4$</u>
(m) nitric acid	<u>HNO_3</u>
(n) carbonic acid	<u>H_2CO_3</u>
(o) sulfuric acid	<u>H_2SO_4</u>
(p) lead(II) nitrate	<u>$Pb(NO_3)_2$</u>
(q) phosphoric acid	<u>H_3PO_4</u>
(r) copper(II) hydroxide	<u>$Cu(OH)_2$</u>
(s) iron(II) phosphate	<u>$Fe_3(PO_4)_2$</u>
(t) calcium chlorate	<u>$Ca(ClO_3)_2$</u>

2. Write the names for the following compounds

2. Write the names for the following compounds.

(a) Li_2CO_3 lithium carbonate

(b) AlHCO_3 aluminum bicarbonate

(c) $\text{Mg}_3(\text{PO}_4)_2$ magnesium phosphate

(d) $\text{Ca}(\text{NO}_3)_2$ calcium nitrate

(e) K_2SO_4 potassium sulfate

(f) $\text{HNO}_3(\text{aq})$ nitric acid

(g) NaNO_3 sodium nitrate

¹⁺(h) $\text{Al}(\text{OH})_3$ aluminum hydroxide

²⁺ ²⁻(i) CuSO_4 copper(II) sulfate

²⁺ ³⁻(j) $\text{Fe}(\text{ClO})_3$ iron (III) hypochlorite

²⁺ ³⁺ ³⁻ (k) $\text{Pb}_3(\text{PO}_4)_2$ lead (II) phosphate

(l) $\text{Sn}(\text{ClO}_3)_2$	tin (II) chlorate
(m) NaOH	sodium hydroxide
(n) $\text{H}_3\text{PO}_4(\text{aq})$	phosphoric acid
(o) $\text{H}_2\text{CO}_3(\text{aq})$	carbonic acid
(p) CuNO_3	copper (I) nitrate
(q) $\text{H}_2\text{SO}_4(\text{aq})$	sulfuric acid
(r) FeSO_4	iron (II) sulfate
(s) $\text{Ca}(\text{HCO}_3)_2$	calcium bicarbonate hydrogen carbonate
(t) K_3PO_4	potassium phosphate

Review Ionic Compounds

Regular

Write Name:

LiCl lithium chloride
 Mg₃S₂ magnesium sulfide

Write Formula:

$Na^{+1} S^{2-}$
 sodium sulfide Na₂S
 potassium bromide KBr

Multivalent

$3^{+} 2^{+} 2^{-}$
 FeO iron(II) oxide
 $1^{+} 2^{+}$ CuCl₂ copper(II) chloride

$K^{+} Br^{-1}$
 $Mn^{4+} O^{2-}$
 manganese (IV) oxide MnO₂
 $V^{5+} N^{3-}$
 vanadium (V) nitride V₃N₅

Polyatomic

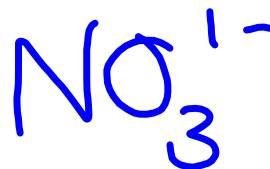
Mg₂(SO₄)₃ magnesium sulfate
Ca(OH)₂ calcium hydroxide

$Ca^{2+} PO_4^{3-}$
 calcium phosphate Ca₃(PO₄)₂
 beryllium sulfate BeSO₄

nitride



nitrate



Review WS ALL Ionic Compounds