

May 24, 2019

Answers pg 241 #1-3
Review Reactions Types

Test Tuesday on Chp 6!!

Warm- Up

For each equation state the reaction type:

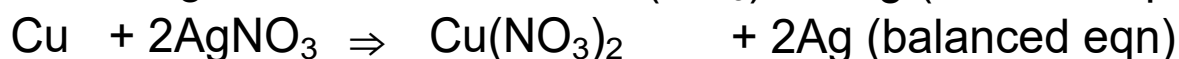
Combustion (C) , Synthesis (S), Decomposition (D), Single Replacement (SR), double replacement (DR)

- $\overset{e}{Ba} + \overset{c}{H_3PO_4} \Rightarrow \overset{c}{Ba_3(PO_4)_2} + \overset{e}{H_2}$
- sodium + bromine \Rightarrow sodium bromide
- methane + oxygen \Rightarrow carbon dioxide + water vapour
- $\overset{c}{CaCl_2} + \overset{c}{Al_2(SO_4)_3} \Rightarrow \overset{c}{CaSO_4} + \overset{c}{AlCl_3}$

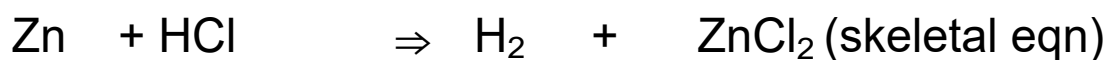
Answers pg 241 #1-3

- two compounds as reactants = double replacement
 - one element and one compound as reactants = single displacement
 - two elements as reactants = synthesis
 - one compound as a reactant = decomposition
- copper + silver nitrate \longrightarrow silver + copper (II) nitrate
single displacement
 - zinc + hydrochloric acid \longrightarrow hydrogen + zinc chloride
single displacement
 - calcium carbonate + hydrochloric acid \longrightarrow carbonic acid + calcium chloride
double replacement
 - aluminum + copper (II) chloride \longrightarrow copper + aluminum chloride
single replacement

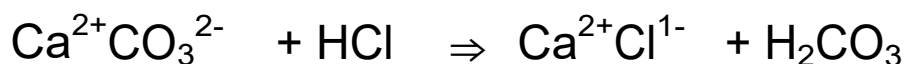
3. a) copper + silver nitrate \Rightarrow silver + copper (II) nitrate



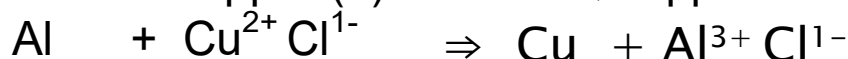
b) zinc + hydrochloric acid \Rightarrow hydrogen + zinc chloride



c) calcium carbonate + hydrochloric acid \Rightarrow carbonic acid + calcium chloride



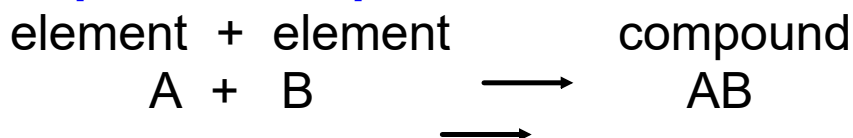
d) aluminum + copper (II) chloride \Rightarrow copper + aluminum chloride



Types of Reactions

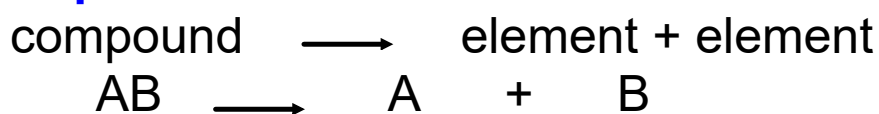
Synthesis (Building Up)

One compound as a product

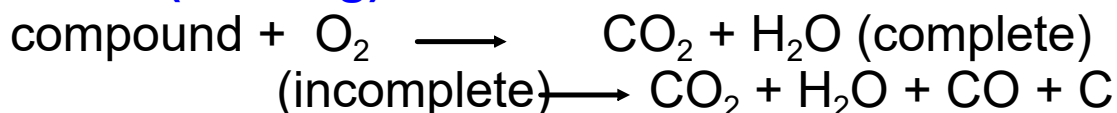


Decomposition (Breaking Down)

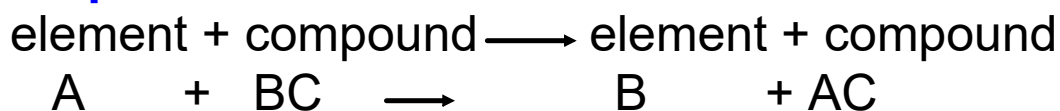
One compound as a reactant



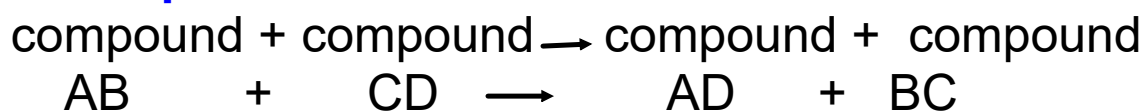
Combustion (burning)



Single Replacement



Double Replacement



Homework

Reactions Worksheet

pg 247 #4,5