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| **CBLT1: Explain the formation, structure and properties of ionic and metallic compounds.** | | |
| Be able to define, explain, identify or provide examples of each of the following: | | |
| * Valence Electrons * Core Electrons * Octet Rule * Cation * Anion * Alloy | * Electron Dot Diagram * Salts * Noble Gas Configuration * Pseudo Noble Gas Config. * Formula Unit * Crystalline Structure | * Metallic Structure * Coordination Number * Electron Sea * Malleable * Ductile |
| Textbook Practice   * Page 193 #s 3 – 11 * Page 199 #s 14 – 17, 20 – 22 * Page 203 #s 23, 25 – 28 | * Page 207 #s 30 – 40, 43, 44, 46 – 49 * Page 208 – 209 #s 53, 55, 56, 58, 59, 60, 67, 70, 73, 76, 78, 79, 81 | |

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| **CBLT2: Define, explain and model molecules, molecular compounds and the different types of covalent bonds with electron dot and structural diagrams.** | | |
| Be able to define, explain, identify or provide examples of each of the following: | | |
| * Covalent Bond * Molecule * Diatomic Molecule * Molecular Compound * Molecular Formula | * Noble Gas Configuration * Single Covalent Bond * Structural Formula * Unshared Pair | * Double Covalent Bond * Triple Covalent Bond * Coordinate Covalent Bond * Polyatomic Ion |
| Textbook Practice – For each question asking for the dot structure, also draw the structural diagram (with lines) | | |
| * Page 216 #s 1 – 3, 6 * Page 220 #s 7, 8 * Page 225 #s 9 – 12 | * Page 229 #s 13 – 16, 20, 21 * Page 247 – 249 #s 39 – 47, 63, 64, 70a, 73, 79, 80 | |

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| **CBLT3: Define, explain, identify and apply sigma and pi bonds and molecular shapes using VSEPR theory.** | | |
| Be able to define, explain, identify or provide examples of each of the following: | | |
| * Molecular Orbitals * Bonding Orbital | * Sigma Bond * Pi Bond | * Tetrahedral * VSEPR |
| Textbook Practice | | |
| * Page 236 #s 23, 24, 27, 29 | * Page 247 – 249 #s 53, 54, 65, 68, 75 | |

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| **CBLT4: Define, explain, identify and apply polar and nonpolar bonds and the different types of intermolecular forces.** | | |
| Be able to define, explain, identify or provide examples of each of the following: | | |
| * Nonpolar Covalent Bond * Polar Covalent Bond * Polar Bond * Polar Molecule | * Dipole * Van der Waals Forces * Dipole Interactions | * Dispersion Forces * Hydrogen Bonds * Network Solids |
| Textbook Practice | | |
| * Page 239 #s 30, 31 * Page 244 #s 32 – 38 | * Page 247 – 249 #s 57 – 61, 70, 72, 74 (research structures) * Also be able to answer, “At room temperature and pressure, why is fluorine and chlorine a gas, bromine a liquid and iodine a solid?” | |