

Oct 25, 2019

go over answers pg 247 #4,5/ Reaction Worksheet
more Chp 6 Review

**Test Monday on Chp 6 Chemical
Reactions!!**

Answers pg 247 #4,5

4. a) synthesis

b) single replacement

c) double replacement

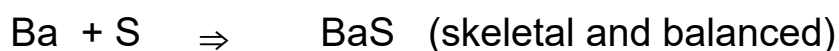
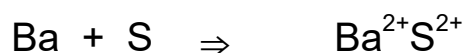
d) decomposition

e) decomposition

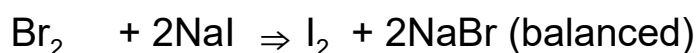
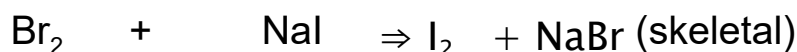
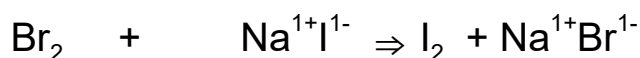
f) single replacement

g) synthesis

5. a) barium + sulfur \Rightarrow barium sulfide



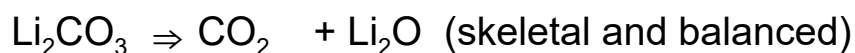
b) bromine + sodium iodide \Rightarrow iodine + sodium bromide



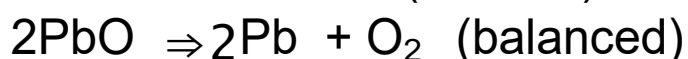
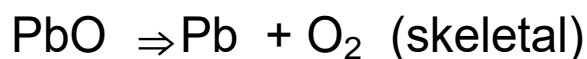
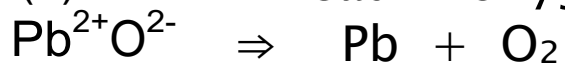
c) barium nitrate + sodium sulfide \Rightarrow barium sulfide + sodium nitrate



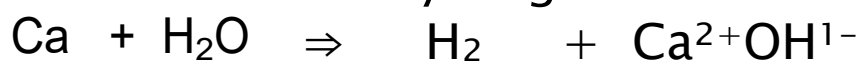
d) lithium carbonate \Rightarrow carbon dioxide + lithium oxide



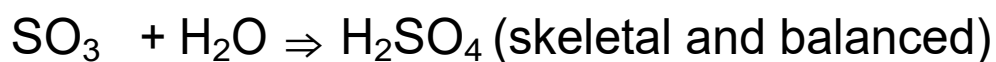
e) lead (II) oxide \Rightarrow lead + oxygen



f) calcium + water \Rightarrow hydrogen + calcium hydroxide



g) sulfur trioxide + water \Rightarrow sulfuric acid



- | | | |
|-----|--|--|
| 1) | $2 \text{NaBr} + 1 \text{Ca(OH)}_2 \rightarrow 1 \text{CaBr}_2 + 2 \text{NaOH}$ | Type of reaction: double displacement |
| 2) | $2 \text{NH}_3 + 1 \text{H}_2\text{SO}_4 \rightarrow 1 (\text{NH}_4)_2\text{SO}_4$ | Type of reaction: synthesis |
| 3) | $4 \text{C}_5\text{H}_9\text{O} + 27 \text{O}_2 \rightarrow 20 \text{CO}_2 + 18 \text{H}_2\text{O}$ | Type of reaction: combustion |
| 4) | $3 \text{Pb} + 2 \text{H}_3\text{PO}_4 \rightarrow 3 \text{H}_2 + 1 \text{Pb}_3(\text{PO}_4)_2$ | Type of reaction: single replacement |
| 5) | $1 \text{Li}_3\text{N} + 3 \text{NH}_4\text{NO}_3 \rightarrow 3 \text{LiNO}_3 + 1 (\text{NH}_4)_3\text{N}$ | Type of reaction: double replacement |
| 6) | $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ | Type of reaction : decomposition |
| 7) | $2 \text{KBr} + \text{F}_2 \rightarrow 2 \text{KF} + \text{Br}_2$ | Type of reaction: single replacement |
| 8) | $\text{Na}_3\text{PO}_4 + 3 \text{KOH} \rightarrow 3 \text{NaOH} + \text{K}_3\text{PO}_4$ | Type of reaction: double replacement |
| 9) | $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{LiCl}$ | Type of reaction: double replacement |
| 10) | $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$ | Type of reaction: decomposition |
| 11) | $2 \text{C}_5\text{H}_5 + \text{Fe} \rightarrow \text{Fe}(\text{C}_5\text{H}_5)_2$ | Type of reaction: synthesis |
| 12) | $\text{SeCl}_6 + \text{O}_2 \rightarrow \text{SeO}_2 + 3\text{Cl}_2$ | Type of reaction: single replacement |
| 13) | $\text{C}_3\text{H}_6\text{O} + 4 \text{O}_2 \rightarrow 3 \text{CO}_2 + 3 \text{H}_2\text{O}$ | Type of reaction: combustion |
| 14) | $2 \text{NO}_2 \rightarrow 2 \text{O}_2 + \text{N}_2$ | Type of reaction : decomposition |
| 15) | $1 \text{AlCl}_3 + 3 \text{Cs} \rightarrow 3 \text{CsCl} + 1 \text{Al}$ | Type of reaction: Single Displacement |

Types of Reactions

Synthesis (Building Up)

One product

Decomposition (Breaking Down)

One Reactant

Combustion (burning)

complete products are $\text{CO}_2 + \text{H}_2\text{O}$

incomplete products are $\text{CO}_2 + \text{CO} + \text{C} + \text{H}_2\text{O}$

Single Replacement

element + compound as reactants and products

Double Replacement

compound + compound as reactants and products

Remember these tips when balancing

- * Remember you can only add co-efficient's
- * Balance the easy atoms first (those that only appear once on each side of the equation)
- * Keep poly-atomics that stay together together.
- * Keep oxygen till the end
- * if there is an odd number of oxygen balance the other atoms 1st then put the odd number in front of the O_2 and double the remaining coefficients.
- * if you have an OH on one side and an H_2O on the other rewrite the H_2O as HOH.

Complete

Chp 6 Review WS/Types of Reaction Review WS

Practice, Practice, Practice these are the same types of questions that are on your test!!!!

Answers will be posted on the website make sure you check!!!

Also make sure you can identify reaction type, write skeletal equations and balance!!! (There is a question on the test like this worth a lot of marks!!!!)