

**Sept 25, 2019**

Answers Review ALL Ionic Compounds Worksheet  
Quiz  
Molecular Compounds

**Quiz Today!!**  
**Test Tuesday!!**



## Name each of the following compounds:

- |   |   |
|---|---|
| 1. $\text{NH}_4\text{Cl}$ - ammonium chloride           | 11. $\text{Sc}(\text{OH})_3$ - scandium hydroxide       |
| 2. $\text{TiBr}_3$ - titanium (III) bromide             | 12. $\text{V}_2(\text{SO}_4)_5$ - vanadium (V) sulphate |
| 3. $\text{Cu}_3\text{P}$ - copper (I) phosphide         | 13. $\text{NH}_4\text{F}$ - ammonium fluoride           |
| 4. $\text{Sn}_3\text{Se}_2$ - tin (II) selenide         | 14. $\text{Li}_2\text{SO}_4$ - lithium sulfate          |
| 5. $\text{Pb}(\text{SO}_4)_2$ - lead (IV) sulfate       | 15. $\text{Zn}_3\text{P}_2$ - zinc phosphide            |
| 6. $\text{RbHCO}_3$ - rubidium bicarbonate              | 16. $\text{Cu}_2\text{O}$ - copper (I) oxide            |
| 7. $\text{NiPO}_4$ - nickel (III) phosphate             | 17. $\text{Ag}_3\text{PO}_4$ - silver phosphate         |
| 8. $\text{MnSO}_4$ - manganese (IV) sulfate             | 18. $\text{CoCO}_3$ - cobalt (II) carbonate             |
| 9. $\text{GaAs}$ - gallium <del>arsenide</del> arsenide | 19. $\text{Fe}(\text{HCO}_3)_2$ - iron (II) bicarbonate |
| 10. $\text{NaBr}$ - sodium bromide                      | 20. $\text{HgO}$ - mercury (II) oxide                   |

## Quiz Ionic Compounds

- all answers go on in the space provided
- make sure you put your name on it

## MOLECULAR COMPOUNDS

-made up of two or more non-metals.

-do not form ions in order to bond



-atoms involved in molecular compounds share electrons.

-the bonds formed by the sharing of electrons are called covalent bonds. A covalent bond is a pair of shared electrons.

-see Fig. 3 page 202

-the diagram of the two chlorines is an example of a diatomic molecule. (Two atoms of the same element). This happens mainly with elements in Table 1-page 202.

$\text{CO}_2$  Carbon dioxide

1 H Hydrogen																	2 He Helium										
3 Li Lithium	4 Be Beryllium											5 B Boron	6 C Carbon	7 N Nitrogen	8 O Oxygen	9 F Fluorine	10 Ne Neon										
11 Na Sodium	12 Mg Magnesium											13 Al Aluminium	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon										
19 K Potassium	20 Ca Calcium	21 Sc Scandium	22 Ti Titanium	23 V Vanadium	24 Cr Chromium	25 Mn Manganese	26 Fe Iron	27 Co Cobalt	28 Ni Nickel	29 Cu Copper	30 Zn Zinc	31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton										
37 Rb Rubidium	38 Sr Strontium	39 Y Yttrium	40 Zr Zirconium	41 Nb Niobium	42 Mo Molybdenum	43 Tc Technetium	44 Ru Ruthenium	45 Rh Rhodium	46 Pd Palladium	47 Ag Silver	48 Cd Cadmium	49 In Indium	50 Sn Tin	51 Sb Antimony	52 Te Tellurium	53 I Iodine	54 Xe Xenon										
55 Cs Cesium	56 Ba Barium	57 La Lanthanum	72 Hf Hafnium	73 Ta Tantalum	74 W Tungsten	75 Re Rhenium	76 Os Osmium	77 Ir Iridium	78 Pt Platinum	79 Au Gold	80 Hg Mercury	81 Tl Thallium	82 Pb Lead	83 Bi Bismuth	84 Po Polonium	85 At Astatine	86 Rn Radon										
87 Fr Francium	88 Ra Radium	89 Ac Actinium	104 Rf Rutherfordium	105 Db Dubnium	106 Sg Seaborgium	107 Bh Bohrium	108 Hs Hassium	109 Mt Meitnerium																			
														58 Ce Cerium	59 Pr Praseodymium	60 Nd Neodymium	61 Pm Promethium	62 Sm Samarium	63 Eu Europium	64 Gd Gadolinium	65 Tb Terbium	66 Dy Dysprosium	67 Ho Holmium	68 Er Erbium	69 Tm Thulium	70 Yb Ytterbium	71 Lu Lutetium
														90 Th Thorium	91 Pa Protactinium	92 U Uranium	93 Np Neptunium	94 Pu Plutonium	95 Am Americium	96 Cm Curium	97 Bk Berkelium	98 Cf Californium	99 Es Einsteinium	100 Fm Fermium	101 Md Mendelevium	102 No Nobelium	103 Lr Lawrencium

Periodic table

## Naming Molecular Compounds

- named similarly to ionic compounds (remember though molecular compounds are two non-metals)



Steps

- Name first element listed, with a prefix to count number of atoms  
 \*do not use a prefix for one atom of the first element\*
- Name second element, with a prefix to count number of atoms.
- Change the suffix of second non-metal to -ide

# of Atoms	Prefix
1	mono
2	di
3	tri
4	tetra
5	penta

