Sept 25, 2019

Answers Review ALL Ionic Compounds Worksheet
Quiz
Molecular Compounds

Quiz Today!!
Test Tuesday!!

Apr 20-11:43 AM

Answers Extra Practice ALL Ionic Compounds

1. Aluminum nitrate
   Al^{3+} NO_3^{1-}
   $Al(NO_3)_3$
2. Calcium fluoride
   Ca^{2+} F^{1-}
   CaF_2
3. Sodium bromide
   Na^{1+} Br^{1-}
   $NaBr$
4. Cobalt (III) chloride
   Co^{3+} Cl^{1-}
   $CoCl_3$
5. Magnesium chlorate
   Mg^{2+} ClO_3^{1-}
   $Mg(ClO_3)_2$
6. Potassium hydroxide
   K^{1+} OH^{1-}
   KOH
7. Nickel (II) oxide
   Ni^{2+} O^{2-}
   NiO
8. Lead (IV) selenide
   Pb^{4+} Se^{2-}
   Pb_2Se_4 $PbSe_2$
9. Cobalt (II) nitrate
   Co^{2+} NO_3^{1-}
   $Co(NO_3)_2$
10. Vanadium (V) phosphate
    V^{5+} PO_4^{3-}
    $V_3(PO_4)_5$
11. Magnesium astatide
    Mg^{2+} At^{1-}
    $MgAt_2$
12. Lithium nitride
    Li^{1+} N^{-3}
    Li_3N
13. Barium nitrate
    Ba^{2+} NO_3^{1-}
    $Ba(NO_3)_2$
14. Aluminum hydroxide
    Al^{3+} OH^{1-}
    $Al(OH)_3$
15. Gallium sulfide
    Ga^{3+} S^{2-}
    Ga_2S_3
16. Chromium (III) bromide
    Cr^{3+} Br^{1-}
    $CrBr_3$
17. Aluminum oxide
    Al^{3+} O^{2-}
    Al_2O_3
18. Zinc sulfate
    Zn^{2+} SO_4^{2-}
    $ZnSO_4$
19. Silver bicarbonate
    Ag^{1+} HCO_3^{1-}
    $AgHCO_3$
20. Gold (I) oxide
    Au^{1+} O^{2-}
    Au_2O

Dec 7-4:06 PM

Name each of the following compounds:

1. NH_4Cl - ammonium chloride
2. $TiBr_3$ - titanium (III) bromide
3. Cu_3P - copper (I) phosphide
4. Sn_3Se_2 - tin (II) selenide
5. $Pb(SO_4)_2$ - lead (IV) sulfate
6. $RbHCO_3$ - rubidium bicarbonate
7. $NiPO_4$ - nickel (III) phosphate
8. $MnSO_4$ - manganese (IV) sulfate
9. GaAs – gallium ~~astatide~~ arsenide
10. NaBr – sodium bromide
11. $Sc(OH)_3$ - scandium hydroxide
12. $V_2(SO_4)_5$ - vanadium (V) sulphate
13. NH_4F – ammonium flouride
14. Li_2SO_4 - lithium sulfate
15. Zn_3P_2 - zinc phosphide
16. Cu_2O – copper (I) oxide
17. Ag_3PO_4 - silver phosphate
18. $CoCO_3$ - cobalt (II) carbonate
19. $Fe(HCO_3)_2$ - iron (II) bicarbonate
20. HgO – mercury (II) oxide

Dec 8-9:45 AM

Quiz Ionic Compounds

- all answers go on in the space provided
- make sure you put your name on it

May 9-8:19 AM

MOLECULAR COMPOUNDS

-made up of two or more non-metals.

-do not form ions in order to bond

-atoms involved in molecular compounds share electrons.

-the bonds formed by the sharing of electrons are called covalent bonds. A covalent bond is a pair of shared electrons.

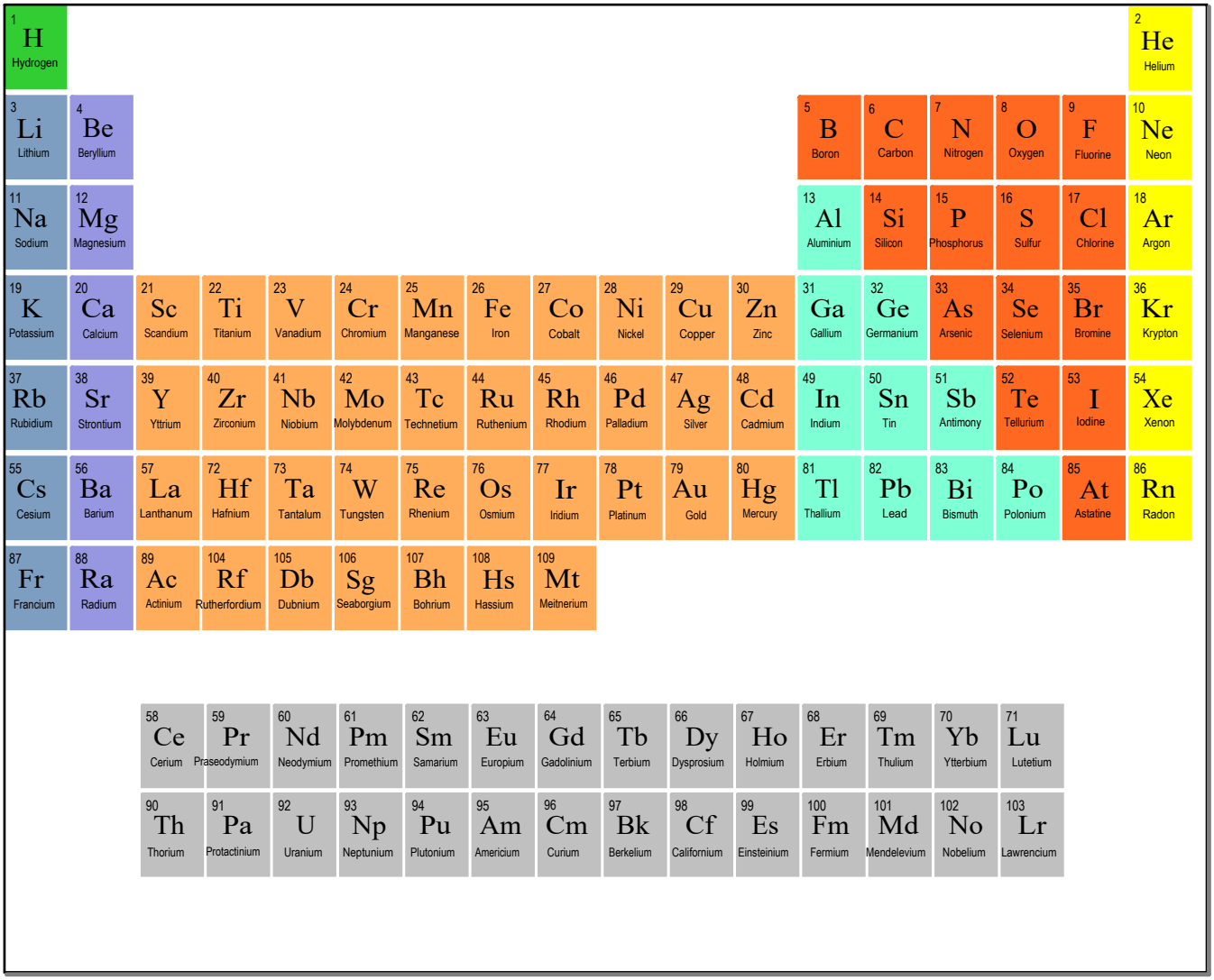
-see Fig. 3 page 202

-the diagram of the two chlorines is an example of a diatomic molecule. (Two atoms of the same element). This happens mainly with elements in Table 1-page 202.

Sep 25-6:26 PM

CO_2 Carbon dioxide

Sep 25-1:20 PM



Periodic table

Naming Molecular Compounds

- named similarly to ionic compounds (remember though molecular compounds are two non-metals) CO_2 carbon dioxide

Steps

1. Name first element listed, with a prefix to count number of atoms (**do not use a prefix for one atom of the first element**)
2. Name second element, with a prefix to count number of atoms.
3. Change the suffix of second non-metal to -ide

# of Atoms	Prefix
1	mono
2	di
3	tri
4	tetra
5	penta

H_2O dihydrogen monoxide

Oct 29-1:57 PM