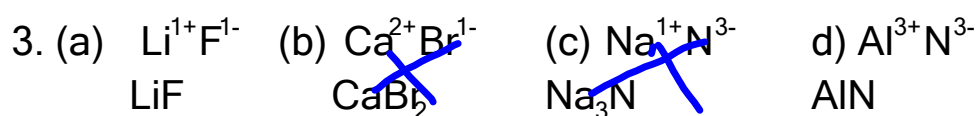


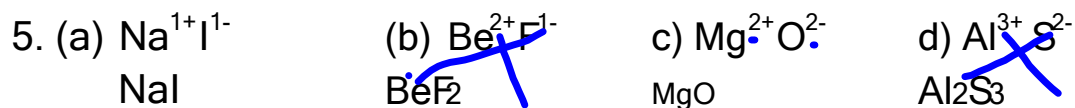
Sept 17, 2019

Answers pg 195 #3-6/WS
Quiz

pg 195 #3-6



4. lithium fluoride calcium bromide sodium nitride aluminum nitride



6. (a) potassium chloride
(b) sodium phosphide
(c) calcium fluoride

1. Lithium and bromine

Name: lithium bromide

Formula:



8. Cesium and arsenic

Name: cesium arsenide

Formula:



2. Strontium and oxygen

Name: strontium oxide

Formula:



9. Beryllium and phosphorus

Name: beryllium phosphide

Formula:



3. Calcium and fluorine

Name: calcium fluoride

Formula:



10. Germanium and oxygen

Name: germanium oxide

Formula:



4. Scandium and chlorine

Name: scandium chloride

Formula:



11. Aluminum and nitrogen

Name: aluminum nitride

Formula:



5. Sodium and sulphur

Name: sodium sulphide

Formula:



12. Barium and chlorine

Name: barium chloride

Formula:



6. Gallium and oxygen

Name: gallium oxide

Formula:



13. Lithium and sulphur

Name: lithium sulphide

Formula:



7. Potassium and bromine

Name: potassium bromide

Formula:



14. Silver and phosphorus

Name: silver phosphide

Formula:



pg 195 #7

Ionic Compounds Quiz

I II III IV V

Sodium oxide
magnesium oxide
manganese II oxide

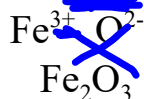
Multi-Valent Ions

some metals have more than one charge they are called **multi-valent ions**

- these elements are found in the middle block of the periodic table i.e. Fe, Ni, Sn, Hg, Cu, Au etc (Table 2 p. 195)

When asked to write the formula the charge that is to be used is indicated in brackets with a Roman numeral

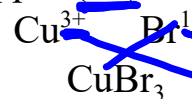
Ex. iron(III) oxide



copper (II) chloride



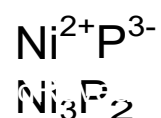
copper (III) bromide



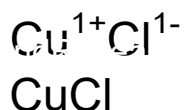
Try These

Write the formula for each of the following:

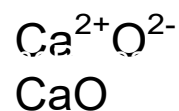
1. nickel (II) phosphide



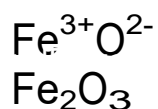
2. copper (I) chloride



3. calcium oxide



4. Iron (III) oxide



5. sodium sulfide

