

# April 24, 2019

answers pg 195 #8-10

Multivalent Ionic Compounds extra practice

## Warm-Up

Name each of the following:

1. MgO magnesium oxide

2. NiO <sup>2+</sup><sub>3+</sub> <sup>2-</sup> nickel (III) oxide

3. PbO<sub>2</sub> <sup>4+</sup><sub>2+</sub> <sup>2-</sup> lead (IV) oxide

Write formulas for each of the following:

1. gold (III) bromide Au<sup>3+</sup> Br<sup>1-</sup>

2. barium oxide Ba<sup>2+</sup> O<sup>2-</sup>

3. cobalt (III) nitride

Co<sup>3+</sup> N<sup>3-</sup>  
CoN

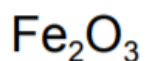
BaO

8a)  $\text{SnCl}_2$  tin (II) chloride

b)  $\text{SnCl}_4$  tin (IV) chloride

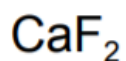
c)  $\text{PbBr}_2$  lead (II) bromide

9a)  $\text{Fe}^{3+} \text{O}^{2-}$



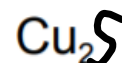
iron (III) oxide

b)  $\text{Ca}^{2+} \text{F}^{1-}$



calcium fluoride

c)  $\text{Cu}^{1+} \text{S}^{2-}$



copper (I) sulphide

10. Iron or Fe

# Ionic Cmpds WS