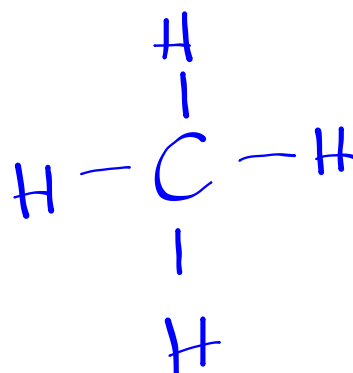
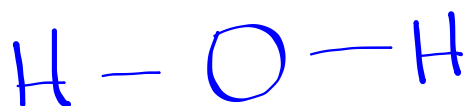


## **Unit 2 - Compounds**

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- Properties of Ionic Compounds, Molecular Compounds, Acids, and Bases (Empirical and Theoretical)
- Naming Ionic Compounds
- Writing formulas for Ionic Compounds
- Ionic hydrates
- Naming Molecular Compounds
- Writing formulas for Molecular Compounds
- Molecular Elements
- Drawing structural diagrams
- Naming and writing formulas for Acids and Bases
- Lab - Identifying Unknown Compounds

14 15 16 17 18  
4 3 2 1 x



$P_4, S_8$

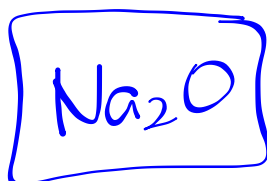
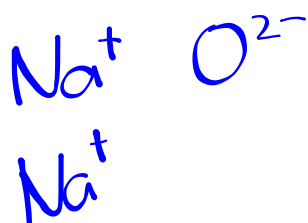
Br HONClIF

$O_2 N_2$



calcium chloride

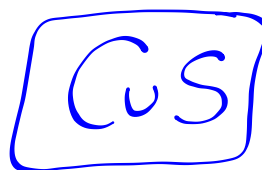
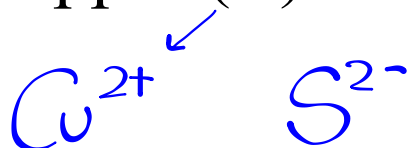
sodium oxide





iron(III) chloride

copper (II) sulfide

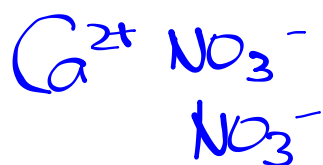
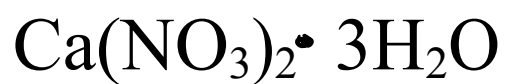




lithium nitrate

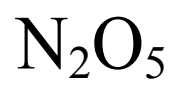
sodium sulfate





Calcium nitrate trihydrate

Calcium nitrate - 3 - water



dinitrogen pentoxide



HClO

$H^+ ClO^-$  aqueous hydrogen hypochlorite  
hypochlorous acid

<sup>-ate</sup>  
chromic acid

$H^+ CrO_4^{2-}$

$H^+$

$H_2CrO_4$

**Review Questions p. 281-282**

**Worksheets**

**# 53, 57-61, 65-71**