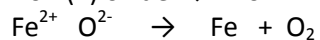


Answers Physical Science 10
Exam Review#2: Chemical Reactions

2.

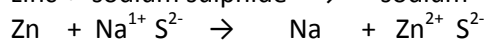
i. iron (II) oxide → iron + oxygen



Reaction Type = Decomposition



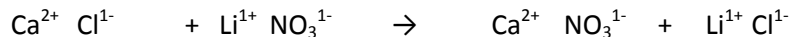
ii. zinc + sodium sulphide → sodium + zinc sulfide



Reaction Type = Single Replacement



iii. calcium chloride + lithium nitrate → calcium nitrate + lithium chloride



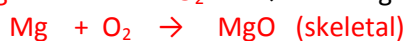
Reaction Type = Double Replacement



iv. magnesium + oxygen → magnesium oxide



Reaction Type = Synthesis



v. chlorine + silver bromide → bromine + silver chloride



Reaction type = Single Replacement



vi. copper + barium sulphate → barium + copper (I) sulphate



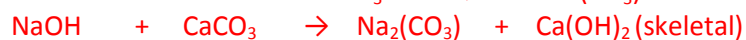
Reaction Type: Single Replacement



vii. sodium hydroxide + calcium carbonate → sodium carbonate + calcium hydroxide



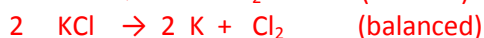
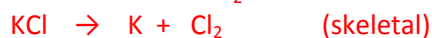
Reaction type = Double Replacement



viii. potassium chloride → potassium + chlorine



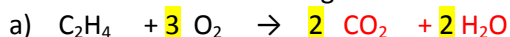
Reaction Type = Decompos



ix. magnesium nitrate + sodium hydroxide → magnesium hydroxide + sodium nitrate



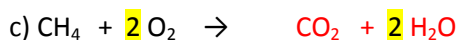
3. For each of the following combustion reactions, complete the balance chemical reaction



complete combustion



incomplete combustion



complete combustion

4. Solution A has a mass of 62g. Solution B has a mass of 89g. When they are mixed, a chemical reaction occurs in which a gas is produced. If the mass of the final mixture is 146g, what mass of gas was produced? Define the law of conservation of mass using this example.

A = 62g

B = 89g

A + B → product + gas

62g + 89g → 146g + gas

151g → 146g + ?

151g – 146g = ?

5g = gas

the law of conservation of mass states that the mass of the reactants must equal the mass of the products. The mass that is missing is of the gas that has evaporated during the reaction.